

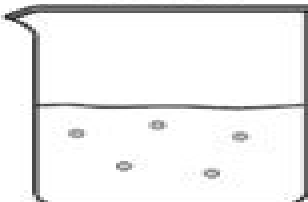
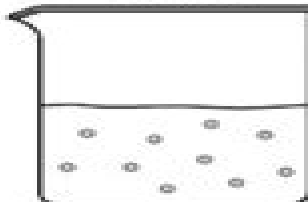
Saturated and Unsaturated Solutions

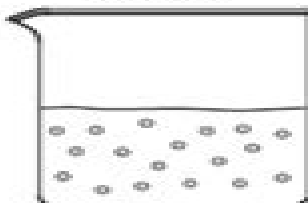
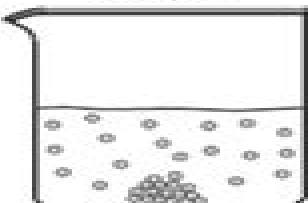
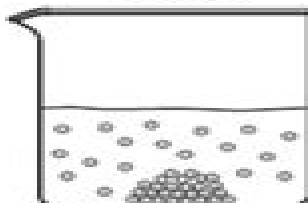
Is there a limit to the amount of solute that will dissolve in a solvent?

Why?

We use solutions every day. People who wear contact lenses use "lens solution" to rinse their contacts and keep them wet. Athletes who consume sports drinks after exercising benefit from the electrolytes in those solutions. This activity will explore whether or not there is a limit to how much of one substance can dissolve in another.

Model 1 – Saturated and Unsaturated Solutions

- All beakers contain 10.0 g of water. - All beakers are kept at 20 °C. - All solutions are stirred for 2 hours. - Solute is the same substance in all beakers.	Unsaturated Solutions	
	Beaker A  1.0 g of solute added	Beaker B  2.0 g of solute added
	Number of dissolved particles	5
	Number of solid particles	0

Saturated Solutions		
Beaker C  3.6 g of solute added	Beaker D  7.0 g of solute added	Beaker E  9.0 g of solute added
		18
	17	

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Rules Of Solubility Pogil Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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