

Fill in the gaps in statements and/or equations.

1. The theory that explains the behaviour of gases on the molecular level is called the kinetic theory, which is based on assumptions related to theoretical gas behaviour, namely that an air particle moves.
2. Use the kinetic theory to explain why a balloon filled with air expands when it is brought inside a warmer room from the cold outdoors.
 According to the kinetic theory, particles have more kinetic energy. Consequently, the particles move faster and exert more pressure on the walls of the balloon, which causes it to expand to a larger volume (until the internal pressure is equal to the external pressure).
3. Define "1 bar" and "1 Pa".
 1 bar is defined as 100,000 Pa. 1 Pa is defined as 1 N/m².
4. State whether each of the listed quantities is an intensive or extensive property.
Temperature and mass are extensive.
5. Pressure is defined as force per unit area. The earth's atmosphere has weight, which creates atmospheric pressure.
6. Standard temperature has been established as 0 °C or 273 K. Standard pressure has been established as 101 kPa, 101 kPa, 101 kPa, or 101 kPa. The atmosphere has a standard temperature and pressure of 101 kPa.
7. A barometer is used to measure atmospheric pressure, which is atmospheric pressure or atmospheric pressure.
8. With atmospheric pressure increasing, how does the height of the column of mercury change? It increases.
9. At constant pressure, the volume of a sample of gas increases with temperature increases.
10. At constant pressure, the volume of a sample of gas is inversely proportional to a temperature decrease with the temperature decrease.

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