

Chapter 6: Chemical Bonds Review Answer Key

- What is the most stable group on the Periodic Table? Why is this group stable?
Noble Gases (Group 18) because they have 8 valence electrons which means their outer energy level is full.
- Name the most reactive group(s) on the Periodic Table? Why are they reactive?
Alkali Metals (Group 1) & Halogens (Group 17): Alkali Metals want to lose one valence electron in order to become stable and Halogens want to gain one valence electron to be stable.

- In the boxes below draw the Electron dot diagrams for:

Sodium	Calcium	Boron	Chlorine	Krypton
$\text{Na} \cdot$	$\text{Ca} \cdot$ $\cdot$	$\cdot \text{B} \cdot$ $\cdot$	$\begin{array}{c} \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \end{array}$	$\begin{array}{c} \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \end{array}$

- Name the three (3) types of chemical bonds:
Ionic bonds Covalent bonds Polyatomic
- Define ionic bond:
The force that holds cations and anions together.
- Define covalent bond:
A chemical bond in which two atoms share a pair of valence electrons.
- Define polyatomic ion:
A covalently bonded group of atoms that has a positive or negative charge and acts as a unit.
- In an ionic bond, electrons are transferred.
- Who are the participants in an ionic bond? Metals and nonmetals
- An ion with a positive charge is a cation; an ion with a negative charge is an anion.
- In a covalent bond, electrons are shared.
- Who are the participants in a covalent bond? Two nonmetals
- Identify whether the following compounds are covalent (two nonmetals), ionic (a nonmetal and a metal), or polyatomic (both).
 - CaCl_2 IONIC (transferred)
 - H_2O COVALENT (shared)
 - BaSO_4 POLYATOMIC (both)
 - CH_4 COVALENT (shared)
 - LiBr IONIC (transferred)
 - NH_4Cl POLYATOMIC (both)
 - KI IONIC (transferred)

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Table of Contents Physical Science Chapter 6 Chemical Bonds Wordwise Answers

1. Understanding the eBook Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - The Rise of Digital Reading Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - Personalized Recommendations
 - Physical Science Chapter 6 Chemical Bonds Wordwise Answers User Reviews and Ratings
 - Physical Science Chapter 6 Chemical Bonds Wordwise Answers and Bestseller Lists
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- ePub, PDF, MOBI, and More
 - Physical Science Chapter 6 Chemical Bonds Wordwise Answers Compatibility with Devices
 - Physical Science Chapter 6 Chemical Bonds Wordwise Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - Highlighting and Note-Taking Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - Interactive Elements Physical Science Chapter 6 Chemical Bonds Wordwise Answers
8. Staying Engaged with Physical Science Chapter 6 Chemical Bonds Wordwise Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 6 Chemical Bonds Wordwise Answers
9. Balancing eBooks and Physical Books Physical Science Chapter 6 Chemical Bonds Wordwise Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 6 Chemical Bonds Wordwise Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter 6 Chemical Bonds Wordwise Answers
- Setting Reading Goals Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter 6 Chemical Bonds Wordwise Answers
- Fact-Checking eBook Content of Physical Science Chapter 6 Chemical Bonds Wordwise Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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