REDUCTION-OXIDATION REACTIONS

Q1 - Identifying Oxidation and Reduction

Assign exidation numbers to each element in the following equations then identify:

- oxidation and reduction
- oxidising and reducing agents

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- c) H,5 + (NO.) 5 + NO +H,0
- d) Cd + NIO + H,O + Cd(OH), + NI(OH),

Q2 - Balancing by Half Equations

Balance each of the following redax reactions using the half equation method.

- a) $(Cr_2O_7)^2 + CT \rightarrow Cr^3 + Cl_2$ (acidic) $(Cr_2O_7)^2 + 6Cl^2 + 14H^2 = 2Cl^3 + 3Cl_3 + 7H_2O$
- b) Cu + (NO₁) → Cu² + NO₂ (ecidic) Cu + 2(NO₂) + 4H → Cu²⁺ + 2NO + ZH₂O
- c) (NO) ALAS TSH) AL(OH) (OH) (NO2) + 2AI + 5H2O + (OH) => 2AI(OH)4- + NH3
- d) (MnO4) + CH2OH -> Mn2+ + HCO2H (acidic) 4 (RHO4)+ + 50R,00H + 12H + 44H2+ 5400H+140
- e) (NO2) + (Cr2O2)2 + Cr2+ (NO2) (acidic) 3(UO2)+ (C2O2)+ 8H 3 2Cr2+3(NO2)+4H0
- 1) cr(OH) + (ClO) + (CrO4)2 + Cl2 (basic) 2 (r(CH)) + 6 (CD) = 2 (CrO4)2 + 8CL+24,0+2(OH)
- g) Iz + (CIO) + (IO) + CI (acidic) Iz + 5(CIO) + +10- 2(D) +5CIT +24+
- h) AsiO1 + (NO1) + H1ASO4 + N2O1 (ocidic) AsiO1 + 2(NO) + 2H2O+2H+ + 2H2HO2 + N2O1
- i) H2Oz + ClOz + (ClOz) + Oz (basic) H0Oz + 2ClOz + 2(OA) 2(ClOz) + Oz + 2H2O
- 1) (Cr2Or)2 . I -> Cr2 . (IO2) (scidic) (CrO2)2 . I + 8H1 = 2Cr2 + (ID2) . 9H30
- k) (MnO1) + Cl + Mn2 + Cl2 (acidic) 2 (HnO1) + 10Cl , 16Ht = 2Mn2 + 5Cl2+ 8 100
- 1) (MnO) + Br + MnO2 + (BrO2) (basic) 2(MnO2) + Br + HO + 2MrO2 + (BrO2) + 2(OH)
- m) As + (ClO2) H2ASO4 + HClO (acidic) 4 A . 5(00) 6 HO. 5H 4 HAQ + 5HCIO
- n) (Cr2O2)2 + CH3OH -> Cr3 + HCO2H (ocidic) 4(Cr2O2)2, 9(H3OH -20H2 -> BC22 + 9HO3H -19H3O
- o) Pb(OH)+2 + (ClO) → PbO+ + CT (basic) Pb (OH)+ + (ClO) PbQ+CI+H+0+2(OH)-
- p) Mn3" + NaBiO3 -> Bi3" + (MnO4)" + Na" (acidic) 2 Hn2" + 5 Not; O3 + (Hn1 55: 3" + 2 (HnO4)"
- q) $H_2O_2 + Cl_2O_7 \rightarrow (ClO_2)^2 + O_2$ (basic)

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Oxidation Reduction Concept Review Answers

Leo J. Malone

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