

## Oxidation – Reduction Packet



Redox reactions occur whenever an electron is transferred from one substance to another. In a redox reaction, one substance is oxidized and another is reduced.

**Oxidation:** the loss of electrons. In a redox reaction, the substance that is oxidized loses electrons. For example, when zinc metal reacts with copper(II) ions, the zinc metal is oxidized to zinc ions, losing two electrons.

**Reduction:** the gain of electrons. In a redox reaction, the substance that is reduced gains electrons. For example, when zinc metal reacts with copper(II) ions, the copper(II) ions are reduced to copper metal, gaining two electrons.

Remember, for every electron lost by one substance, another substance gains the same number of electrons. This is the law of conservation of charge.



Remember, the total number of electrons lost must equal the total number of electrons gained.

The two half-reactions are coupled together, meaning the electrons lost by one substance are gained by another. This is why the total number of electrons lost must equal the total number of electrons gained. This is why the total number of electrons lost and gained must be equal.

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