#### Chemistry -- Gas Laws ACT Questions [USE TO STUDY FOR YOUR TEST]

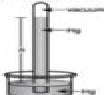
#### Multiple Choice

Identify the choice that best completes the statement or answers the question.

At 22°C, the air pressure in a car tire is 1293 torr.

Convert this pressure to atmospheres (atm).

- a. 12.76
- b. 5.262
- c. 1.701
- d. 1.293
- 2. This diagram shows a closed-end mercury (Hg)



barometer.

When atmospheric pressure is 1.30 atm, what is the height (h), in millimeters, of the Hg in the barometer?

- a. 132
- b. 585
- c. 760
- d. 988
- A chemist compresses a 1.5 L sample of carbon dioxide (CO<sub>2</sub>) gas at 1.2 atm to a final volume of 0.75 L. Assuming the temperature is constant, what is the final pressure, in atm, exerted by the CO<sub>2</sub> gas?
  - a. 0.60
  - b. 0.94
  - c. 1.7
  - d. 2.4
- A vessel connected to an open-end mercury (Hg)
  manometer contains nitrogen (N<sub>2</sub>) gas. The
  atmospheric pressure is 0.993 atm. The height
  difference between the two arms of the manometer



is 122 mm.

What is the pressure, in mm Hg, of the N2 gas?

- a. 355
- b. 633

- c. 755
- d. 877
- At 1.00 atm, a scaled weather balloon contains 20.0 L of helium (He) gas at 25.0°C. Assume that none of the He escapes and the pressure is constant. What is the volume, in liters, of He in the weather balloon at 35.0°C?
  - a. 40.0
  - b. 28.0
  - c. 20.7
  - d. 14.3
- 6. At 1254 m above sea level, the city of Helena, Montana, has an average atmospheric pressure of 0.859 atm. A chemistry student in Helena finds that pure distilled water boils at 95.9°C. Which statement explains the student's results at this high elevation?
  - The boiling point of water increases due to lower atmospheric pressure.
  - The boiling point of water increases due to higher atmospheric pressure.
  - The boiling point of water decreases due to lower atmospheric pressure.
  - The boiling point of water decreases due to higher atmospheric pressure.
- For a chemistry homework assignment, Beth must consider the effusion rates of methane (CH<sub>4</sub>) gas and nitrogen trifluoride (NF<sub>2</sub>) gas at 25°C. Effusion occurs as gas particles escape from a container through a small hole. Which statement accurately compares the effusion rates of CH<sub>4</sub> and NF<sub>3</sub> at 25°C?
  - a. CH<sub>4</sub> effuses 2 times faster than NF<sub>3</sub>.
  - b. CH<sub>4</sub> effuses 4 times faster than NF<sub>3</sub>.
  - c. NF3 effuses 2 times faster than CH4.
  - d. NF<sub>3</sub> effuses 4 times faster than CH<sub>4</sub>.
- At standard temperature and pessure, a chemist burns 3.50 L of pentane (C<sub>2</sub>H<sub>12</sub>) in excess oxygen (O<sub>2</sub>) gas. C<sub>2</sub>H<sub>12</sub>(I) + 8 O<sub>2</sub>(g) ⇒ 5 CO<sub>2</sub>(g) + 6 H<sub>2</sub>O(I) What is the maximum number of liters of carbon dioxide (CO<sub>2</sub>) that the reaction can produce?
  - a. 17.5
  - b. 8,50
  - c. 5.60
  - d. 0.700

# **Sample Test Questions On Gas Laws**

**Yi-Tong Ma** 

**Sample Test Questions On Gas Laws:** 

This book delves into Sample Test Questions On Gas Laws. Sample Test Questions On Gas Laws is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Sample Test Questions On Gas Laws, encompassing both the fundamentals and more intricate discussions.

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