

Chapter 12 Review 1: Covalent Bonds and Molecular Structure

1) How are ionic bonds and covalent bonds different, and what types of elements combine to form each?

Ionic bonds result from the transfer of electrons from one atom to another (formed by a metal and a non-metal).

Covalent bonds result from two atoms sharing electrons (formed by 2 or more non-metals).

2) How are nonpolar covalent bonds different from covalent bonds, and what types of elements combine to form each?

polar – electrons are shared unequally – between nonmetals with different electronegativities

nonpolar – electrons are shared equally – between multiple atoms of the same element

3) Identify the type(s) of bond(s) found in the following molecules:

- a. CCl_4 _____ **covalent** _____
- b. Li_2O _____ **ionic** _____
- c. NF_3 _____ **covalent** _____
- d. CaSO_4 _____ **ionic and covalent** _____
- e. SO_2 _____ **covalent** _____
- f. Mg(OH)_2 _____ **ionic and covalent** _____

4) Define electronegativity.

Electronegativity is the ability for an atom in a molecule to attract electrons to itself.

5) Use electronegativity values to place the following elements in increasing order: F, N, Si, C, O

$\text{Si} < \text{C} < \text{N} < \text{O} < \text{F}$

6) Determine if the bond between atoms in each example below is nonpolar covalent, polar covalent, or ionic.

- | | | | |
|-------------------|--------------------------|------------------|--------------------------|
| a. H_2 | _____ npv _____ | e. NF_3 | _____ pc _____ |
| b. PCl_3 | _____ pc _____ | f. MgO | _____ ionic _____ |
| c. F_2 | _____ npv _____ | g. CH_4 | _____ npv _____ |
| d. NaBr | _____ ionic _____ | h. HCl | _____ pc _____ |

Reviewing Key Concepts Chapter 12

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