

CHAPTER 9 REVIEW

Stoichiometry

KEY

SECTION 3

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

$$\% \text{ yield} = \frac{22 \text{ g}}{25 \text{ g}} \times 100 = 88\%$$

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation:

2.0 mol x 3 N_2

ratio

6.0 mol N_2 12.0 mol H_2

- a. Which chemical is in excess? What is the excess in moles?

$$\text{N}_2: \left(\frac{2 \text{ mol N}_2}{1 \text{ H}_2} \right) \left(\frac{1 \text{ H}_2}{3 \text{ H}_2} \right) = 9.0 \text{ mol excess}$$

not $3 \times \text{N}_2 \Rightarrow \text{H}_2$ is limiting8.0 mol NH_3

- b. Theoretically, how many moles of NH_3 will be produced?

$$\text{mol NH}_3 = \left(\frac{12.0 \text{ mol H}_2}{1 \text{ H}_2} \right) \left(\frac{2 \text{ NH}_3}{3 \text{ H}_2} \right) = 8.0 \text{ mol}$$

6.4 mol NH_3

- c. If the percentage yield of NH_3 is 80%, how many moles of NH_3 are actually produced?

$$\% \text{ yield} = \frac{\text{actual}}{\text{theor}} \times 100$$

$$\text{actual} = \frac{(\% \text{ yield})(\text{theor})}{100} = \frac{(80)(8.0)}{100} = 6.4 \text{ mol}$$

3. 0.050 mol of $\text{Ca}(\text{OH})_2$ are combined with 0.080 mol of HCl according to the following equation:

0.10 mol HCl

- a. How many moles of HCl are required to neutralize all 0.050 mol of $\text{Ca}(\text{OH})_2$?

$$\text{mol HCl} = \left(\frac{0.050 \text{ mol Ca}(\text{OH})_2}{1 \text{ Ca}(\text{OH})_2} \right) \left(\frac{2 \text{ HCl}}{1 \text{ Ca}(\text{OH})_2} \right) = 0.10 \text{ mol HCl}$$

limiting

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Table of Contents Review Stoichiometry Section 2 Answers Modern Chemistry

1. Understanding the eBook Review Stoichiometry Section 2 Answers Modern Chemistry
 - The Rise of Digital Reading Review Stoichiometry Section 2 Answers Modern Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Stoichiometry Section 2 Answers Modern Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Stoichiometry Section 2 Answers Modern Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Stoichiometry Section 2 Answers Modern Chemistry

- Personalized Recommendations
- Review Stoichiometry Section 2 Answers Modern Chemistry User Reviews and Ratings
- Review Stoichiometry Section 2 Answers Modern Chemistry and Bestseller Lists
- 5. Accessing Review Stoichiometry Section 2 Answers Modern Chemistry Free and Paid eBooks
 - Review Stoichiometry Section 2 Answers Modern Chemistry Public Domain eBooks
 - Review Stoichiometry Section 2 Answers Modern Chemistry eBook Subscription Services
 - Review Stoichiometry Section 2 Answers Modern Chemistry Budget-Friendly Options
- 6. Navigating Review Stoichiometry Section 2 Answers Modern Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
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 - Review Stoichiometry Section 2 Answers Modern Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Stoichiometry Section 2 Answers Modern Chemistry
 - Highlighting and Note-Taking Review Stoichiometry Section 2 Answers Modern Chemistry
 - Interactive Elements Review Stoichiometry Section 2 Answers Modern Chemistry
- 8. Staying Engaged with Review Stoichiometry Section 2 Answers Modern Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Stoichiometry Section 2 Answers Modern Chemistry
- 9. Balancing eBooks and Physical Books Review Stoichiometry Section 2 Answers Modern Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Stoichiometry Section 2 Answers Modern Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Stoichiometry Section 2 Answers Modern Chemistry
 - Setting Reading Goals Review Stoichiometry Section 2 Answers Modern Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Stoichiometry Section 2 Answers Modern Chemistry

- Fact-Checking eBook Content of Review Stoichiometry Section 2 Answers Modern Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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