

Review Stoichiometry Chemistry

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- How many moles of tungsten atoms are in 4.8×10^{25} atoms of tungsten?
A 8.0×10^2 moles
B 8.0×10^1 moles
C 1.3×10^{-1} moles
D 1.3×10^{-2} moles
- What is the molar mass of $(\text{NH}_4)_2\text{CO}_3$?
A 144 g
B 138 g
C 96 g
D 78 g
- What is the number of moles of beryllium atoms in 36 g of Be?
A 0.25 mol
B 4.0 mol
C 45.0 mol
D 320 mol
- In every chemical reaction, _____.
A mass and molecules are conserved
B moles and liters are conserved
C mass and atoms are conserved
D moles and molecules are conserved
- In the reaction $2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g)$, what is the ratio of oxygen used to moles of CO_2 produced?
A 1:1
B 2:1
C 1:2
D 2:2
- How many moles of aluminum are needed to react completely with 1.2 mol of FeO ?
 $2\text{Al}(s) + 3\text{FeO}(s) \rightarrow 3\text{Fe}(s) + \text{Al}_2\text{O}_3(s)$
A 1.2 mol
B 0.8 mol
C 1.6 mol
D 2.4 mol
- The equation below shows the decomposition of lead nitrate. How many grams of oxygen are produced when 11.5 g NO_2 is formed?
 $2\text{Pb}(\text{NO}_3)_2(s) \rightarrow 2\text{PbO}(s) + 4\text{NO}_2(g) + \text{O}_2(g)$
A 1.00 g
B 2.00 g
C 2.88 g
D 32.0 g
- How many grams of H_3PO_4 are produced when 10.0 moles of water react with an excess of P_4O_{10} ?
 $\text{P}_4\text{O}_{10}(s) + 6\text{H}_2\text{O}(l) \rightarrow 4\text{H}_3\text{PO}_4(aq)$
A 1.22 g
B 6.7 g
C 147 g
D 653 g
- Which element has a molar mass of 30.974 g/mol?
A Potassium
B Phosphorus
C Gallium
D Palladium
- Which is the mass of 8 moles of sodium chloride?
A 7.3 grams
B 468 grams
C 0.137 g/mole
D 468 moles
- How many molecules are in 3.6 grams of NaCl ?
A 0.06
B 1.0×10^{26}
C 1.3×10^{24}
D 3.7×10^{21}
- How many grams are in 1.946 moles of NaCl ?
A 113.8 g
B 30.1 g
C 0.033 g
D 44.7 g

Review Stoichiometry Modern Chemistry Answers

Y Pai

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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