

Chapter 9 – Stoichiometry – Chapter Review

Last page worked out for you.

Match each description on the right to the correct term on the left.

- | | | |
|----------|--|----------------------|
| <u>B</u> | 1. The substance that determines the amount of product that can be formed in a reaction. | A. Percent yield |
| <u>E</u> | 2. The amount of product that forms when a reaction is carried out in the laboratory. | B. Limiting reagent |
| <u>D</u> | 3. The calculation of quantities in chemical equations. | C. Theoretical yield |
| <u>A</u> | 4. The ratio of the actual yield to the theoretical yield expressed as a percent. | D. Stoichiometry |
| <u>F</u> | 5. The substance that is present in enough quantity to react with a limiting reagent. | E. Actual yield |
| <u>C</u> | 6. The maximum amount of products that could be formed from given amounts of reactants. | F. Excess reagent |

Multiple Choice - Choose the best answer and write its letter in the blank.

- D 7. How many moles of H_2O would be required to produce 2.5 moles of O_2 ?
- $$2 \text{H}_2\text{O} \rightarrow 2 \text{H}_2 + \text{O}_2$$
- A. 2.0 mol C. 4.0 mol
B. 2.5 mol D. 5.0 mol
- D 8. If 3.00 mol of CaCO_3 undergo decomposition to form CaO and CO_2 , how many grams of CO_2 are produced?
- $$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$$
- A. 3.00 g C. 88.0 g
B. 44.0 g D. 132 g
- B 9. How many moles of H_2O are produced when 240 grams of CuO react?
- $$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$$
- A. 1.0 mol C. 18 mol
B. 3.0 mol D. 54 mol

Review Vocabulary Ideal Amount Of Product Answer Key Stoichiometry

F. Vögtle, E. Weber



Review Vocabulary Ideal Amount Of Product Answer Key Stoichiometry:

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