Skill - Student can use a balanced equation to determine mole to mole ratios.

- 1) Given this equation, N. + 3 H. + 2 NH., write the following molar ratios.

  - ber Marc Meta Brain
  - CONTRACT MARKS THE PARTY.
- Given the following equation: 8 H<sub>2</sub> + S<sub>2</sub> → 8 H<sub>2</sub>S, write the following molar ratios:
  - A H. H. H. C. C. C.
  - by Mary Co. But II.
  - CONTRACTOR OF THE REAL PROPERTY.
- 30 Answer the following questions for this equation: 2 H. + O. + 2 Hub
  - What is the H<sub>a</sub>: H<sub>a</sub>O molar ratio?
  - b. With 20 moles of H<sub>a</sub> on hand & plenty of O<sub>a</sub>, how many moles of H<sub>a</sub>O could you make?
  - c. What is the Out Hub molar ratio?
  - d. With 29 moles of 0<sub>3</sub> and enough H<sub>a</sub>, how many moles of H<sub>2</sub>O could you make?
- 40 American the following questions for this equation: N. + 3 M. + 2 NH.
  - a. If you used 1 mole of N<sub>s</sub>, how many males of NH<sub>s</sub> could be produced?
  - if 50 moles of fift, were produced, how many moles of N, would be required?
     10mol 50H, v (1)mol N, (2mol 50H, ) = 5 mol N.
  - c. If 3.00 moles of Humere used, how many moles of NH, would be made?
  - d. if 0.600 moles of NH, were produced, how many moles of M, are required? 0.600 mol NH, x (Simol H,/2 mol NH,) = 0.9 mol H.

Skill - Student can use a balanced equation, given the grams of reactant, to calculate the orange of another substance

5) Magnesium reacts with hydrochloric acid according to the following equation:

He fet + 2 HCl fact + He/I, tack + H, foll

# 72.92 grams of hydrochiloric acid react with excess magnesium, how many moles of hydrogen gas will be produced?

72.92g HCl x 1mml HCl x 1.mml H, = 1.000weel H, 35.46gHCl 2mml HCl

- 6) Aluminum reacts with hydrobromic acid to produce aluminum bromide and hydrogen gas. Write a balanced equation for the reaction and calculate the number of moles of hydrobromic acid required to react with 23.47 grams of aluminum.
  2 Alia) + 6 Http://eq.) 2 Althylio) + 3 H.(q)
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**Michael Seilmaier** 

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