



**NATIONAL
SENIOR CERTIFICATE**

GRADE 11

NOVEMBER 2020

**PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR**

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use appropriate units for all calculations.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B? at 0 °C and 101 kPa of external pressure? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its relative value of these molecules. (2)
- 3.1.4 What effect, if any, does compound B have on the normal boiling point of compound A? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 What compound (A or B) will have the higher surface tension? (2)

Compound B (M_r 102) will show in the table, has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound B is 16,2 °C (lower than that of compound A) and its relative value of ΔH_{vap} and ΔH_{fusion} in kJ mol⁻¹. (2)

3.2 Substances C1 is added to water whereas 2 propanone (acetone) and water is separated into two layers in the diagram below.



- 3.2.1 Is 2-propanone (C3H8O) or H₂O, the water dissolved? (2)
- 3.2.2 Explain the process in QUESTION 3.2.1 about the solubility in the intermolecular forces involved. (2)

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Page 10 of 10

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Table of Contents Physical Science Paper2 November 2014

1. Understanding the eBook Physical Science Paper2 November 2014
 - The Rise of Digital Reading Physical Science Paper2 November 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Paper2 November 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Paper2 November 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Paper2 November 2014

- Personalized Recommendations
- Physical Science Paper2 November 2014 User Reviews and Ratings
- Physical Science Paper2 November 2014 and Bestseller Lists
- 5. Accessing Physical Science Paper2 November 2014 Free and Paid eBooks
 - Physical Science Paper2 November 2014 Public Domain eBooks
 - Physical Science Paper2 November 2014 eBook Subscription Services
 - Physical Science Paper2 November 2014 Budget-Friendly Options
- 6. Navigating Physical Science Paper2 November 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Paper2 November 2014 Compatibility with Devices
 - Physical Science Paper2 November 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Paper2 November 2014
 - Highlighting and Note-Taking Physical Science Paper2 November 2014
 - Interactive Elements Physical Science Paper2 November 2014
- 8. Staying Engaged with Physical Science Paper2 November 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Paper2 November 2014
- 9. Balancing eBooks and Physical Books Physical Science Paper2 November 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Paper2 November 2014
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Paper2 November 2014
 - Setting Reading Goals Physical Science Paper2 November 2014
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Paper2 November 2014

- Fact-Checking eBook Content of Physical Science Paper2 November 2014
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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