

Solution Manual for Physical Chemistry 2nd Edition Ball

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4.14.

$$\left(\frac{\partial H_m}{\partial T}\right)_P = \frac{C_{P,m}}{T} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\text{and } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

and that we can estimate $C_{P,m}$ at 298 K as:

- 4.15. (a) The noble gas with the fewest atoms has the smallest percentage of noble gas.
 (b) The noble gas with the greatest number of atoms has the greatest percentage of noble gas.
 (c) The noble gas with the greatest number of atoms has the smallest percentage of noble gas.

4.16. From the data we find the heat capacity of the gas is $C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$, and that the heat capacity is $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$.

$$\text{or } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

4.17. From the data we find $C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$, and that the heat capacity is $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$.

$$\text{and } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

$$\text{and } C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

4.18. From the data we find the heat capacity of the gas is $C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$, and that the heat capacity is $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$.

$$\text{or } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

Physical Chemistry Ball Solution Manual

Julia Schneider



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