

Group 15 Elements

N, P, As, Sb, Bi, Mc

Hydrides of Group 15 Elements

Halides of Group 15 Elements The valence shell electronic configuration is ns<sup>2</sup>np<sup>3</sup>

N and P - non-metals
As and Sb - metalloids
Bi and Mc (moscovium) - metals

Common oxidation states are -3, +3 and +5 NH<sub>3</sub> has higher melting and boiling points than PH<sub>3</sub> due to hydrogen bonding

NH<sub>3</sub> is a mild reducing agent and BiH<sub>3</sub> is the strongest reducing agent

Basicity order - NH<sub>3</sub> > PH<sub>3</sub> > AsH<sub>3</sub> > SbH<sub>3</sub> > BiH<sub>3</sub>

Pentahalides are more covalent than trihalides

Nitrogen does not form pentahalide due to absence of the d-orbitals in its valence shell

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