

Unit 7: Reduction, Oxidation and ElectrochemistryChapter 17: Electrochemistry4.9: Oxidation-Reduction Reactions

Reduction-Oxidation Reactions (Redox Rxn): - chemical reactions where there is a transfer of electron(s).

Oxidation States (Oxidation Number): - a number that is arbitrary assigned to an atom in an element, molecule, and polyatomic ions to account for the number of electrons in *redox* reaction.

Rules for Assigning Oxidation Numbers:

1. Natural Elements (include diatomic and polyatomics) have an Oxidation Number of 0.

Examples: Na_(s), O_{2(g)}, O_{3(g)}, H_{2(g)}, F_{2(g)}, P_{4(s)}, and Hg_(l) all have an oxidation number of 0.

2. Single Atomic Ions have an Oxidation Number Equals to its Charge.

Example: K⁺ has an oxidation number of +1.

3. Oxygen in Binary Compound and Polyatomic Ions has an Oxidation Number of -2.

Examples: CO₂, CO, SO₃, SO₂, CO₃²⁻, SO₄²⁻ and H₂O all have an oxidation number of -2 for oxygen.

An exception occurs in peroxides (compound containing O₂²⁻) where the oxygen has an oxidation number of -1.

4. Hydrogen in Binary Compounds and Polyatomic Ions has an Oxidation Number of +1.

Examples: H₂O, HCl, HBr, CH₄, NH₄⁺ and H₂S all has an oxidation number of + 1 for hydrogen.

5. In Binary Compounds, the Atom that has the most attraction to electrons (HIGHER Electronegativity) is assigned the NEGATIVE Oxidation Number. This Negative Oxidation Number is the SAME as its Ionic Charge.

Examples: HF: F is more electronegative. Thus, F has an oxidation number of -1.

NH₃: N is more electronegative. Thus, N has an oxidation number of -3.

H₂S: S is more electronegative. Thus, S has an oxidation number of -2.

6. The Sum of all Oxidation Numbers is 0 for all Electrically Neutral Compounds.

Examples: CO: Oxidation numbers: O = -2; C = +2 [(+2) + (-2) = 0]

SO₂: Oxidation numbers: O = -2; S = +4 [(+4) + 2(-2) = 0]

7. The Sum of all Oxidation Numbers is Equal to the Overall Charge for all Polyatomic Ions.

Examples: CO₃²⁻: Oxidation numbers: O = -2; C = +4 [(+4) + 3(-2) = -2]

SO₄²⁻: Oxidation numbers: O = -2; S = +6 [(+6) + 4(-2) = -2]

Note: Same Atom from different chemical species might NOT have the Same Oxidation Number.

We write the charges of ionic species as n^+ or n^- . Oxidation numbers are written as $+n$ or $-n$.

Oxidation Reduction Study Guide Answers

Stephen W Leslie



Oxidation Reduction Study Guide Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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