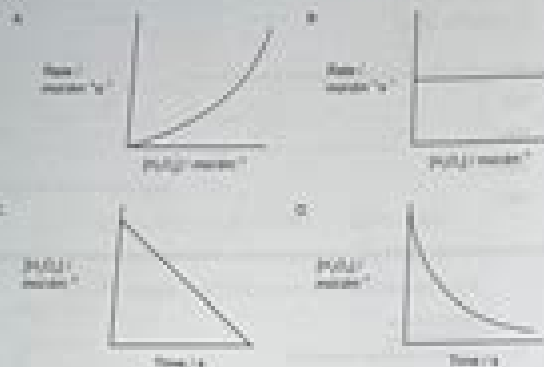


12. Decomposition of hydrogen peroxide in an aqueous solution proceeds as follows:



The rate expression for the reaction was found to be: rate =  $k[\text{H}_2\text{O}_2]$

Which graph is consistent with the given rate expression?



13. The rate constant,  $k$ , is commonly described by the Arrhenius equation:  $k = Ae^{-E_a/RT}$ . Which of the following statements are correct?

- A greater  $E_a$  value results in a smaller  $k$  value.
- Reactions of less complex molecules usually have a greater value of  $A$ .
- The slope (gradient) of a  $\ln k$  versus  $1/T$  equals  $E_a$ .

- A. i and ii only  
B. i and iii only  
C. ii and iii only  
D. i, ii and iii

Turn over

14. What happens when the temperature of the following equilibrium system is increased?



| Position of equilibrium | Equilibrium value of forward and reverse rate constants |
|-------------------------|---------------------------------------------------------|
| A. shifts to the left   | increase                                                |
| B. shifts to the left   | decrease                                                |
| C. shifts to the right  | decrease                                                |
| D. shifts to the right  | increase                                                |

15. A mixture of 5.0 moles of  $\text{CO(g)}$  and 5.0 moles of  $\text{H}_2(\text{g})$  was sealed in a 1.00 dm<sup>3</sup> vessel. The following equilibrium was established:



At equilibrium, the mixture contained 0.20 moles of  $\text{CO(g)}$ . How many moles of  $\text{H}_2(\text{g})$  and  $\text{CH}_3\text{OH(g)}$  were present at equilibrium?

|    | Equilibrium mol of $\text{H}_2$ | Equilibrium mol of $\text{CH}_3\text{OH}$ |
|----|---------------------------------|-------------------------------------------|
| A. | 0.20                            | 0.75                                      |
| B. | 0.50                            | 0.50                                      |
| C. | 0.30                            | 0.20                                      |
| D. | 0.70                            | 0.75                                      |

16. Which species behave as Brønsted-Lowry bases in the following reaction?



- A.  $\text{HNO}_3$  and  $\text{HSO}_4^-$   
B.  $\text{HNO}_3$  and  $\text{H}_2\text{NO}_3^+$   
C.  $\text{H}_2\text{NO}_3^+$  and  $\text{HSO}_4^-$   
D.  $\text{H}_2\text{NO}_3^+$  and  $\text{HNO}_3$

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