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$$= \frac{\text{moles of solute}}{\text{kg of solvent}}$$

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$M_{mix} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$$
 1.5 x 480 + 1.2 x 520 / 480 + 520 = 720 / 1000 = 1.344 M
 with increase in temperature which of these changes a molality b fraction of solute present in water

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