

1. Draw a graph to show the relationship between the number of carbon atoms in each alkane and its boiling point. (Number of carbon atoms will go on the x-axis and boiling point on the y-axis).
2. Describe what you see.
3. Suggest a reason for what you have observed.
4. Why was it enough for us to use 'number of carbon atoms' as a measure of the molecular weight of the molecules?

Click here for the solution⁵

2.5.2 Investigation : Determining the density of liquids:

Density is a very important property because it helps us to identify different materials. Every material, depending on the elements that make it up and the arrangement of its atoms, will have a different density.

The equation for density is:

$$\text{Density} = \frac{\text{Mass}}{\text{Volume}} \quad (2.1)$$

Discussion questions:

To calculate the density of liquids and solids, we need to be able to first determine their mass and volume. As a group, think about the following questions:

- How would you determine the mass of a liquid?
- How would you determine the volume of an irregular solid?

Apparatus:

Laboratory mass balance, 10 ml and 100 ml graduated cylinders, thread, distilled water, two different liquids.

Method:

Determine the density of the distilled water and two liquids as follows:

1. Measure and record the mass of a 10 ml graduated cylinder.
2. Pour an amount of distilled water into the cylinder.
3. Measure and record the combined mass of the water and cylinder.
4. Record the volume of distilled water in the cylinder
5. Empty, clean and dry the graduated cylinder.
6. Repeat the above steps for the other two liquids you have.
7. Complete the table below.

Liquid	Mass (g)	Volume (ml)	Density ($\text{g} \cdot \text{ml}^{-1}$)
Distilled water			
Liquid 1			
Liquid 2			

Table 2.5

2.5.3 Investigation : Determining the density of irregular solids:

Apparatus:

Use the same materials and equipment as before (for the liquids). Also find a number of solids that have an irregular shape.

Method:

Determine the density of irregular solids as follows:

⁵<http://www.flinn.org/HP>

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