

## Worksheet 1-1 - Measuring Reaction Rates

Name - \_\_\_\_\_  
Due Date - \_\_\_\_\_

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

1. A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:



A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

$$\begin{array}{r} 0.016\text{g} \\ - 0.006\text{g} \\ \hline 0.010\text{g} \end{array}$$

- a) Calculate the *Rate of Reaction* in grams of Zn consumed per second.

$$\text{rate} = \frac{\Delta g}{\Delta \text{time}} = \frac{0.010\text{g}}{20\text{sec}}$$

Answer  $5.0 \times 10^{-4}$

- b) Calculate the *Rate of Reaction* in moles of Zn consumed per second.

$$5.0 \times 10^{-4} \frac{\text{g}}{\text{sec}} \times \frac{1\text{mol}}{65.4\text{g}} =$$

Answer  $7.6 \times 10^{-6} \frac{\text{mol}}{\text{sec}}$

- c) Write out the complete ionic equation for the reaction.



- d) What will happen to the  $[\text{H}^+]$  as the reaction proceeds? decrease

- e) What will happen to the  $[\text{Cl}^-]$  as the reaction proceeds? nothing, it is a spectator

# Reaction Rate Worksheet 1 Answers

**H Kauffman**



## **Reaction Rate Worksheet 1 Answers:**

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*Part 503 Implementation Guidance*, 1995

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## **Table of Contents Reaction Rate Worksheet 1 Answers**

1. Understanding the eBook Reaction Rate Worksheet 1 Answers
  - The Rise of Digital Reading Reaction Rate Worksheet 1 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rate Worksheet 1 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reaction Rate Worksheet 1 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rate Worksheet 1 Answers
  - Personalized Recommendations
  - Reaction Rate Worksheet 1 Answers User Reviews and Ratings
  - Reaction Rate Worksheet 1 Answers and Bestseller Lists
5. Accessing Reaction Rate Worksheet 1 Answers Free and Paid eBooks
  - Reaction Rate Worksheet 1 Answers Public Domain eBooks
  - Reaction Rate Worksheet 1 Answers eBook Subscription Services
  - Reaction Rate Worksheet 1 Answers Budget-Friendly Options
6. Navigating Reaction Rate Worksheet 1 Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Reaction Rate Worksheet 1 Answers Compatibility with Devices
  - Reaction Rate Worksheet 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Reaction Rate Worksheet 1 Answers
    - Highlighting and Note-Taking Reaction Rate Worksheet 1 Answers
    - Interactive Elements Reaction Rate Worksheet 1 Answers
  8. Staying Engaged with Reaction Rate Worksheet 1 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Reaction Rate Worksheet 1 Answers
  9. Balancing eBooks and Physical Books Reaction Rate Worksheet 1 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Reaction Rate Worksheet 1 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Reaction Rate Worksheet 1 Answers
    - Setting Reading Goals Reaction Rate Worksheet 1 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Reaction Rate Worksheet 1 Answers
    - Fact-Checking eBook Content of Reaction Rate Worksheet 1 Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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