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Practice: pages 20-25

- 1 Answer is B. From experiments 1 and 3 the order with respect to Cl₂ is 1 (the rate quadruples as the concentration doubles. Therefore, in an experiment using 0.80 mol dm⁻¹ of CO and 0.20 mol dm⁻² of Cl₂, the rate will be halved from the 6.4 × 10⁻⁰⁴ mol dm⁻³ min⁻¹ obtained in experiment 1. (There is no need to work out the order with respect to CO.)
- 2 Answer is B. The concentration halves every 120 s. So, the sequence of half-lives is 2.4 → 1.2 → 0.6 → 0.3 → 0.15; that is four half-lives which is 4 × 120 s. [1]
- Answer is C. For a zero order reaction, the rate of decomposition is a constant. [1]
- 4 Answer is A. the decomposition is Rate = k[X]² [1]
 - From the data given, $0.016 = k[1]^2$ [1]
 - Therefore, $k = 0.016 \text{ s}^{-1}$. If $\{X\} = 0.25 \text{ mol dim}^{-1}$, rate = 0.016 $[0.25]^2$ = 0.001 dm³ mol⁻² s⁻²
- 6 Answer is D. I₂ has an order of zero and is not involved in the rate-determining step. Therefore, it can be ignored. In D. [0.1](2) has a value that is greater than [0.1](1).
- 7 Answer is C. The slow step + the fast step must give the overall equation. I must, therefore, be removed in the fast step in a reaction with G, and some J and K must be formed. [1
- 8 Answer is B. The rate equation is Rate = k[halide] so statements 1 and 2 are correct. Raising the temperature will increase the rate of reaction by allowing more molecules to exceed the activation energy, but the activation energy is fixed and will not change. [1]
- 9 Answer is D. Increasing the amount of ozone present does not change the order with respect to ozone. A large excess is not necessary to ensure enough ozone is present. The purpose of a large excess is to make sure that during the reaction its concentration is virtually constant Therefore, only statement 3 is correct.
- 10 Answer is A. Experiments 2 and 3 give the order with respect to X as zero. Experiments 1 and 2 give the order with respect to Y as 2. So statements 1 and 2 are correct. The rate constant is calculated from a set of results, for example, experiment 1 gives 0.08 = k[0.10]².
- So $k = 8.0 \text{ dm}^3 \text{ mmol}^{-1} \text{ s}^{-1}$ and statement 3 is, therefore, correct. [1]
- 11 a) $2N_2O_3(g) \rightarrow 4NO_2(g) + O_2(g)$ [1]
 - b) i) For every 2 mol of N₂O₃ that reacts 4 mol of NO₂ is formed. Therefore, the rate of formation of NO₃ is twice as fast as the loss of N₂O₃, i.e. 7.0 × 10⁻⁴ mol dm⁻² s⁻¹. [1]
 - For oxygen, the rate of formation will be $\frac{1}{2}$ the rate of loss of N₂O₃. i.e. 1.75×10^{-6} mol dm⁻⁸ s⁻⁶

Reaction Rate And Energy Reinforcement Answers

S Nieto

Reaction Rate And Energy Reinforcement Answers:

Kaplan & Sadock's Study Guide and Self-Examination Review in Psychiatry Benjamin J. Sadock, Virginia A. Sadock, Pedro Ruiz, 2012-03-28 Kaplan Sadock's Study Guide and Self Examination Review in Psychiatry is a comprehensive review of the specialty and perfect for stand alone review or as preparation for the PRITE in service ABPN Part I and recertification examinations The book contains more than 1 600 multiple choice questions and answers with explanatory discussions of correct and incorrect responses Chapters parallel the essential corresponding chapters in Kaplan Sadock's Synopsis of Psychiatry a staple of psychiatry education around the globe Terms and definitions are consistent with DSM IV TR and ICD Behavior Analysis and Learning Erin B. Rasmussen, Casey J. Clay, W. David Pierce, Carl D. Cheney, 2022-12-30 Using a behavioral perspective Behavior Analysis and Learning provides an advanced introduction to the principles of behavior analysis and learned behaviors covering a full range of principles from basic respondent and operant conditioning through applied behavior analysis into cultural design The text uses Darwinian neurophysiological and biological theories and research to inform B F Skinner's philosophy of radical behaviorism The seventh edition expands the focus on neurophysiological mechanisms and their relation to the experimental analysis of behavior providing updated studies and references to reflect current expansions and changes in the field of behavior analysis By bringing together ideas from behavior analysis neuroscience epigenetics and culture under a selectionist framework the text facilitates understanding of behavior at environmental genetic neurophysiological and sociocultural levels This grand synthesis of behavior neuroscience and neurobiology roots behavior firmly in biology The text includes special sections New Directions Focus On Note On On the Applied Side and Advanced Section which enhance student learning and provide greater insight on specific topics This edition was also updated for more inclusive language and representation of people and research across race ethnicity sexuality gender identity and neurodiversity Behavior Analysis and Learning is a valuable resource for advanced undergraduate and graduate students in psychology or other behavior based disciplines especially behavioral neuroscience The text is supported by Support Material that features a robust set of instructor and student resources www routledge com 9781032065144 Energy Research Abstracts ,1990 Behavior Analysis and Learning W. David Pierce, Carl D. Cheney, 2013-04-15 Behavior Analysis and Learning Fourth Edition is an essential textbook covering the basic principles in the field of behavior analysis and learned behaviors as pioneered by B F Skinner The textbook provides an advanced introduction to operant conditioning from a very consistent Skinnerian perspective It covers a range of principles from basic respondent conditioning through applied behavior analysis into cultural design Elaborating on Darwinian components and biological connections with behavior the book treats the topic from a consistent worldview of selectionism The functional relations between the organism and the environment are described and their application in accounting for old behavior and generating new behavior is illustrated Expanding on concepts of past editions the fourth edition provides updated coverage

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