Name: Class:\_ Five Types of Chemical Reaction Worksheet Answers Balance the following reactions and indicate which of the six types of chemical reaction are being 2 NaBr + 1 Ca(OH)<sub>2</sub> → 1 CaBr<sub>2</sub> + 2 NaOH Type of reaction: double displacement að. 2 NH<sub>1</sub>+ 1 H<sub>2</sub>SO<sub>4</sub> → 1 (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> **b**) Type of reaction: synthesis 4 C<sub>5</sub>H<sub>9</sub>O + 27 O<sub>2</sub> → 20 CO<sub>2</sub> + 18 H<sub>2</sub>O Type of reaction: combustion CB. di) 3 Pb + 2 H<sub>2</sub>PO<sub>4</sub> → 3 H<sub>2</sub> + 1 Pb<sub>1</sub>(PO<sub>4</sub>)<sub>2</sub> Type of reaction: single displacement Type of reaction: double displacement et. 1 Li<sub>2</sub>N + 3 NH<sub>4</sub>NO<sub>3</sub> → 3 LiNO<sub>3</sub> + 1 (NH<sub>4</sub>)<sub>3</sub>N Section 3: Predicting the products of chemical reactions 1) 2 Ag + 1 CuSO<sub>4</sub> → 1 Ag<sub>2</sub>SO<sub>4</sub> + 1 Cu Type: Single Displacement 23 2 Nai + 1 CaCl<sub>2</sub> → 2 NaCl + 1 Cal<sub>2</sub> Type: Double Displacement 30 1 O2 + 1 H2 → 2 H2O Type: Synthesis 400 2 AgNO<sub>2</sub> + 1 BaSO<sub>4</sub> → 1 Ag<sub>2</sub>SO<sub>4</sub> + 1 Ba(NO<sub>2</sub>)<sub>2</sub> Type: Double Displacement 51 2 HCN + 1 CuSO<sub>4</sub> → 1 H<sub>2</sub>SO<sub>4</sub> + 1 Cu(CN)<sub>2</sub> Type: Double Displacement 6) 1 H<sub>2</sub>O + 1 AgI → 1 HI + 1 AgOH Type: Double Displacement 73 4 LiBr + 1 Co(SO<sub>1</sub>)<sub>2</sub> → 2 Li<sub>2</sub>SO<sub>1</sub> + 1 CoBr<sub>4</sub> Type: Double Displacement

Type: Single Displacement

Type: Single Displacement

Type: Single Displacement

Type: Double Displacement

Type: Single Displacement

Type: Double Displacement

Type: Decomposition

Type: Synthesis

Type: Decomposition

Type: Combustion

Type: Combustion

8):

9)

1.0%

1.13

12)

13)

14)

151

1.6)

171

18)

1 LiNO<sub>1</sub> + 1 Ag → 1 AgNO<sub>1</sub> + 1 Li

1 N2 + 2 O2 → 2 NO2

1 H<sub>2</sub>CO<sub>3</sub> → 1 CO<sub>2</sub> + 1 H<sub>2</sub>O

1 AlCl<sub>3</sub> + 3 Cs → 3 CsCl + 1 Al

1 C<sub>0</sub>H<sub>0</sub> + 6 O<sub>2</sub> → 4 CO<sub>2</sub> + 4 H<sub>2</sub>O

1 BaS + 1 PtCl<sub>2</sub> → 1 BaCl<sub>2</sub> + 1 PtS

2 Na₂O → 4 Na + 1 O₂

1 Al(NO<sub>3</sub>)<sub>3</sub> + 1 Ga → 1 Ga(NO<sub>3</sub>)<sub>3</sub> + 1 Al

2 KCl + 1 Mg(OH)₂ → 2 KOH + 1 MgCl₂

1 Zn + 1 Au(NO₂)₂ → 1 Zn(NO₂)₂ + 1 Au

1 CH<sub>3</sub>COOH + 2 O<sub>2</sub> → 2 CO<sub>2</sub> + 2 H<sub>2</sub>O

# **Reaction Types Answers Chemistry**

**D** Siedentop

**Reaction Types Answers Chemistry:** 

Unveiling the Magic of Words: A Review of "Reaction Types Answers Chemistry"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is actually aweinspiring. Enter the realm of "**Reaction Types Answers Chemistry**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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