Long answer type

1. i) The rate law is given by, r = [A] 1/2[B] 2. What is the order of this reaction.

ii) A first order reaction is found to have a rate constant $k = 5.5 \times 10-145 - 1$ Find the Half life of the reaction. iii) Write two differences between order and molecularity.

i)A reaction is 50% complete in 20 minutes at 300K and the same reaction again complete in 5 minutes at 350K. Calculate the activation energy if it is first order reaction [R=8.314 JK-1Mole-1 ; log4 = 0.602]. ii)Define rate constant and order of a reaction

3. i)A first order reaction has a rate constant of 0.0051min-1. If we begin with 0.10M concentration of the reactant, what concentration of the reactant will be left after three hour?

ii)A reaction is second order with respect to a reactant .How is its rates effected when concentration of the reactant is

i) doubled ii)reduced to half

 i)Show graphically how the rate of a 1st order reaction with only one reactant depends upon the concentration of the reactant

ii)Decomposition of phosphine (PH3) proceeds according to the following equation 4PH3(g)→P4(g)+6H2(g). It is found that the reaction follows the following rate equation Rate-k (PH3) The half life of (PH3) is 37.9s at1200C.

i)How much time is required for 3/4th of phosphine to decompose.

ii)What fraction of original sample of PH3reemains behind after I minute. Determine the order of the reaction with respect to A and with respect to B and overall order of the reaction.

5(a) A reaction is second order in A and first order in

B. (i). Write the differential rate equation.

(iii). How is the rate affected on increasing the concentrations of A three times?

(iii). How is the rate affected when the concentrations of both A and B are doubled?

(b). A first order reaction takes 40 minutes for 30% decomposition. Calculate Expt. No [A] mole/I [B] mole/I Rate of formation of D mole/I/min 1 0.1 0.1 5.0x10-3 2 0.3 0.2 6.0x10-2 3 0.3 0.4 2.4x10-1 4 0.4 0.1 2.0x10-2 t1/2 for this reaction. (Given log 1.428 = 0.1548)

Reaction Rate Review Answers

Richard Bailey

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