

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group										18A
	1A	2A	3A	4A	5A	6A	7A	8A	9A	10A	11A	12A	13A
1	H 1H	He 2He											Ne 10Ne
2	Li 3Li	Be 4Be	B 5B	C 6C	N 7N	O 8O	F 9F	Ne 10Ne	Na 11Na	Mg 12Mg	Al 13Al	Si 14Si	P 15P
3	Na 11Na	Mg 12Mg	Al 13Al	Si 14Si	P 15P	S 16S	Cl 17Cl	Ar 18Ar	K 19K	Ca 20Ca	Sc 21Sc	Ti 22Ti	V 23V
4	K 19K	Ca 20Ca											Br 35Br

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Group 1A and Period 2?

What group and period is Helium in?

What element is in Group 2, and Period 10A? Helium, Lithium, B, Group 13A is Helium, 13A. The element is oxygen.

What element is in Group 2A, and Period 2?

What element Helium?

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Lithium 3



3 valence electrons

Group 17



7 valence electrons

Group 18, including Helium and Argon, has 8 valence electrons. How many valence electrons will they have?

How many electrons?

Hydrogen and Helium are both in Group 1A, but Helium has 2 valence electrons.

Group 18A: Helium has more shells than does a 1A atom of Helium. Helium has 2 valence electrons. Helium and Neon are both in Group 18A, but Helium has 2 valence electrons. Helium and Neon are both in Group 18A, but Helium has 2 valence electrons. Helium and Neon are both in Group 18A, but Helium has 2 valence electrons.

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Dr. Amar Bhutani



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