

CLO- Classification of Elements & Periodicity in Properties

MULTIPLE CHOICE QUESTIONS:

Choose the one correct option in the following:

- Identify the least stable ion amongst the following :
(a) Li^+ (b) Be^+ (c) B^+ (d) C^+
- The correct order of radii is :
(a) $\text{N} < \text{Be} < \text{B}$ (b) $\text{F}^- < \text{O}^{2-} < \text{N}^{3-}$ (c) $\text{N} < \text{Li} < \text{K}$ (d) $\text{Fe}^{2+} < \text{Fe}^{3+} < \text{Fe}^{4+}$
- Which one of the following has largest size :
(a) Na (b) Na^+ (c) Na^- (d) Li^+
- The number of elements in the 4th period of periodic table is :
(a) 8 (b) 10 (c) 18 (d) 32
- Which of the following has least electron affinity (electron gain enthalpy) :
(a) O (b) N (c) Ar (d) F
- An element of atomic number 29 belongs to which of the following block of the periodic table :
(a) s-block (b) p-block (c) d-block (d) f-block
- Which one of the following arrangements not truly represents the property indicated against it:
(a) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – electronegativity (b) $\text{Br}_2 < \text{F}_2 < \text{Cl}_2$ – Electron Affinity
(c) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – Bond energy (d) $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$ – Oxidising Power
- Alkali metals in each period have :
(a) Smallest size (b) Lowest IE (c) Highest IE (d) Highest EN
- A sudden large jump between the value of second and third ionization energies of elements would be associated with which of the following electronic configuration :
(a) $2\text{p}^4 3\text{s}^1$ (b) $3\text{s}^2 3\text{p}^1$ (c) $3\text{s}^2 3\text{p}^2$ (d) $2\text{p}^4 3\text{s}^2$
- Which of the following pairs are chemically dissimilar :
(a) Na and K (b) Ba and Sr (c) Zr and Hf (d) Ca and Zn
- Collective name given to the elements with outer shell configuration $ns^2 np^6$ is :
(a) Chalkogens (b) Alkaline earths metals (c) Transition metals (d) Nobel gases
- Modern extended periodic table was given by :
(a) Bohr (b) Lothar mayer (c) Moseley (d) None alone
- The most electronegative elements is :
(a) Hydrogen (b) Fluorine (c) Chlorine (d) Oxygen
- The element having electronic configuration :
 $[\text{Kr}]4\text{d}^{10}4\text{f}^{14}5\text{s}^2 5\text{p}^6 5\text{d}^1 6\text{s}^2$ belongs to
(a) s-block (b) p-block (c) d-block (d) f-block
- An elements R forms the highest oxide R_2O_5 . R belong to :

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