

Study Guide: Periodic Table and Periodic Trends (Honors)

1. Differentiate between Mendeleev's periodic table and Moseley's periodic table.

Mendeleev put the elements in order by atomic mass, but he noticed that in several places, the atomic masses did not always increase sequentially. Moseley discovered that ordering by atomic number was the correct way to describe the order.

2. State the periodic law.

The properties of the elements are periodic functions of their atomic numbers.

3. Vertical columns on the periodic table are called **groups or families** and horizontal rows are called **periods**.

4. Elements with similar chemical properties are placed in the same **group**.

5. Complete the table below.

Electron Config.	Block	Period	Group/Type	Element
[Xe] 6s ² 4f ¹⁴ 5d ¹⁰ 6p ⁶	p	6	18 / Noble Gas	Radon
[Xe] 6s ¹	s	6	1 / Alkali Metal	Cesium
[Ar] 4s ¹ 3d ⁵	d	4	6 / Trans. Metal	Chromium
[He] 2s ²	s	2	2 / Alk. Earth Metal	Beryllium
[Kr] 5s ¹ 4d ¹⁰	d	5	11 / Trans. Metal	Silver
[He] 2s ² 2p ⁵	p	2	17 / Halogen	Fluorine

6. Determine whether the following statements are describing alkali metals (A), alkaline earth metals (AE), transition metals (TM), halogens (H), noble gases (NG):

- a) most reactive non-metals **H**
- b) s²p⁶ valence configuration **NG**
- c) second most reactive metals **AE**
- d) s¹ valence configuration **A**
- e) exceptions found to e- configurations **TM**
- f) soft metals **A**
- g) d orbitals being filled **TM**

7. Would you expect Li and S to have different chemical properties? Why? **Yes; Li is an alkali metal with 1 valence e-, and S is a nonmetal with 6 valence e-.**

8. Explain why a magnesium atom is smaller than: a sodium atom; a calcium atom. **Mg is smaller than Na because the greater nuclear charge in Mg pulls the e- cloud in tighter (periodic trend). Mg is smaller than Ca because it has fewer energy levels (group trend).**

9. Predict the size of the astatine atom compared to the sulfur atom. **At is larger than S, because At is 3 periods down from S, but only 1 group over. The trends compete, but the sizeable increase in energy levels overcomes the slight difference in nuclear charge.**

10. Would you expect a Cl⁻ ion to be larger or smaller than an Mg²⁺ ion. Explain. **A Cl⁻ ion is larger than a Mg²⁺ ion, because the chloride ion has 3 energy levels, matching**

Periodic Trends Study Guide

J Ma



Periodic Trends Study Guide:

This book delves into Periodic Trends Study Guide. Periodic Trends Study Guide is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Periodic Trends Study Guide, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Periodic Trends Study Guide
- Chapter 2: Essential Elements of Periodic Trends Study Guide
- Chapter 3: Periodic Trends Study Guide in Everyday Life
- Chapter 4: Periodic Trends Study Guide in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Periodic Trends Study Guide. The first chapter will explore what Periodic Trends Study Guide is, why Periodic Trends Study Guide is vital, and how to effectively learn about Periodic Trends Study Guide.
3. In chapter 2, this book will delve into the foundational concepts of Periodic Trends Study Guide. The second chapter will elucidate the essential principles that need to be understood to grasp Periodic Trends Study Guide in its entirety.
4. In chapter 3, the author will examine the practical applications of Periodic Trends Study Guide in daily life. The third chapter will showcase real-world examples of how Periodic Trends Study Guide can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Periodic Trends Study Guide in specific contexts. The fourth chapter will explore how Periodic Trends Study Guide is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Periodic Trends Study Guide. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Periodic Trends Study Guide.

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