

Annotated periodic table of the elements

1	2	Group names*										13	14	15	16	17	18
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19
3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21
5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22
6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23
7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24

Group names*

- Hydrogen & the alkali metals
- Alkaline earth metals
- Coinage metals
- Volatile metals
- Isotopes
- Crystallogens
- Postcogens
- Chalcogens
- Halogens
- Helium group

*Groups 3 to 10 are named after their first members (i.e. Group 3 is the Scandium group)

Main group
oxidation states
electronegativity
Madelung-Pauling

Name number
and symbol
Name weight
and form

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Coinage
Alkali earth	
Alkali	
Transition	Intermetallic
Post-transition	
Post-transition	Metalloid
Alkali (Fr-Ra, Cs-Ac)	Alkali gas

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Alkali earths**
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. Ca to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Alkalines**
Soft, dense and reactive metals. Those from Ba onwards are synthetic. Early alkalines (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late alkalines are more like lanthanides. Many alkalines form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; inertlike oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metalloids

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in amorphous forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.9–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colorless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metalloids to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

Group 1

- Group 1**
The alkali are Li-F, and the alkaline earths are La-branch-1a - Cs-Ac, and 4 Lu branch (i.e. Lu-Lu). These group two-Sr-Fr - alkalis:
• Lu branch, resemble those of groups 1-2
• Lu branch, more resemble those of groups 4-10
- Coinage nonmetals**
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive
Approximation is likely to be either a metalloid or a post-transition metal

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Daniel F McAuley



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