# **Percent Yield Practice I**

A student uses stoichiometry to <u>calculate</u> the predicted yield of CO<sub>2</sub> in a chemical reaction as <u>323</u> g. When the student *mixes the chemicals* to make the CO<sub>2</sub>, he finds that the reaction only produces <u>308 g</u> of CO<sub>2</sub>. What is the percent yield?

$$Percent \ yield = \frac{actual \ yield}{theoretical \ yield} \times 100$$

Percent yield = 
$$\frac{308}{323} \times 100$$

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