

Orgo 2OG3-Cheat Sheet

	Strong base Weak nucleophile	Strong base Strong nucleophile	Weak base Strong nucleophile	Weak base Weak nucleophile
1°	E2	E2 SN2	SN2	
2°	E2	E2 SN2	SN2	
3°	E2	E2	SN1	SN1 E1

Strong base Weak nucleophile	Strong base Strong nucleophile	Weak base Strong nucleophile	Weak base Weak nucleophile
NaH DBN DBU	HO ⁻ MeO ⁻ EtO ⁻	I ⁻ Br ⁻ Cl ⁻ RS ⁻ HS ⁻ RSH H ₂ S	H ₂ O MeOH EtOH

-SN1 & E1 only seen for 1° & 2° C's when carbocation is resonance stabilized

Predicting Products:

1. Substrate
-1° C, 2° C, 3° C
2. Reagent
-Nucleophilicity/Basicity
3. Stereochemistry
SN2 → Inversion
SN1 → Racemic
E1/E2 → Trans > Cis
4. Regioselectivity
SN1/SN2 → α-carbon
E2 → Zaitsev favoured
-Unless base is hindered
E1 → Zaitsev favoured always

E2: Base deprotonates, halide leaves & π-bond forms
-More exposed adjacent C protons the better
3° C > 2° C > 1° C

Regioselectivity

- Greater substituted alkene increases stability
- Increasing bulky nucleophiles prefer less substituted
- Stereoselectivity
- Trans alkenes > stability than cis alkenes
- Stereospecificity
- Adjacent proton must be anti to leaving group
- Opposite stereochemistry (Dash-Wedge)

SN2: Nucleophile attacks electrophile → Dissociates leaving group
Nucleophile strength → Charge > Basicity > Polarizability
-Back side attack therefore inverse configuration about C atom (S → R)
-Steric bulk prohibits attack
1° C > 2° C > 3° C
-Bulky adjacent carbons disallow interaction (Ex. Neopentyl)
-Bulky nucleophiles also not favoured for SN2
Effect of solvents → Refer loose-leaf

SN1: Alkyl halides undergo ionization in a polar solvent
-LG leaves, nucleophile attacks carbocation
-Carbocation intermediate (Unimolecular)
-Carbocation rearrangement
-Methyl/hydride shifts to stabilize the charge
Stereochemistry
-Attack can occur from two equal paths
-Forms a racemic mixture (Opposite configurations)

E1: After ionization of alkyl halide, deprotonation occurs & π-bond forms
-Carbocation intermediate (Unimolecular)

Regioselectivity

- Greater substituted alkene increases stability
- Trans alkenes > stability than cis alkenes

Hydrohalogenation:

- Addition of H-X across a nucleophilic π-bond
- Markovnikov's Rule:
-H adds to carbon with most H's (Carbocation stability)
- Anti-Markovnikov Addition:
-In Presence of peroxides → Radical's form and H adds to carbon with less H's
- Stereochemistry
-If chiral center formed → Both enantiomers formed (Racemic)

Acid-Catalyzed Hydration (H₂O Addition):

- Addition of H₂O across a double bond in presence of dilute acid



Hydrohalogenation & Hydration Both Involve a Carbocation:

- Have chance of undergoing rearrangement



- If we don't want the rearrangement products
-Oxymercuration & Demercuration

- Follows Markovnikov's rule

Rate of reaction:

3° C > 2° C > 1° C (Carbocation stability)

Alcohol Dehydration:

- Alcohols can be dehydrated using conc. H₂SO₄



Orgo Mechanisms Cheat Sheet

John T. Moore, Richard H. Langley



Orgo Mechanisms Cheat Sheet:

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