

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Covalent Bonds

There are two types of covalent bond: single bond and double bond.

- **Single bond** is formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end overlap).



- **Double bond** is formed by sharing of electron between two atoms.



The Relative Strength of σ and π Bonds



Single and Double Bond



Relative strength



There are two types of covalent bond



Sigma (σ) and pi (π) bonds

Covalent bonds are classified into σ and π bonds.

- **Sigma bond** is formed by head-on overlap of orbitals.

- **Pi bond** is formed by side-on overlap of orbitals.

- **Double bond** is a σ bond.

- **Triple bond** is a σ bond and two π bonds.



σ bond is stronger than π bond.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds are

- **polar** and **nonpolar**.

• If the difference in electronegativity is less than 0.5, the bond is nonpolar.

• If the difference in electronegativity is between 0.5 and 1.7, the bond is polar.

• A difference of 1.7 or higher is ionic.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

- The more electronegative atom has a partial negative charge (δ^-), and the less electronegative atom has a partial positive charge (δ^+).

- δ^- (delta minus) indicates an electron deficient atom.

- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments of the bonds.



Sum of the vector is the net dipole moment of the molecule.

Bond Length and Bond Strength

The length of the bond is measured in Å.

Bond	H-H	H-Cl	H-Br	H-I
Length	0.074 Å	0.127 Å	0.141 Å	0.154 Å
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

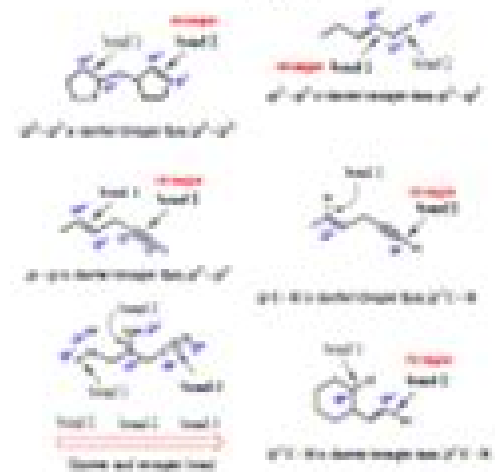


The length of the bond is measured in Å.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	376 kJ/mol	611 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
C-H Bond	1.09 Å	1.06 Å	1.03 Å
Strength	413 kJ/mol	505 kJ/mol	535 kJ/mol

Higher s character, shorter and stronger bond

Examples



giving information on the strength of the bond.



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