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Name Chapter 1—Structure and Bonding
Description
Instructions

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Question 1 **Essay** **0 points**
Question Give the ground-state electron configuration for carbon (atomic number 6).
Answer $1s^2 2s^2 2p_x^1 2p_y^1$ or $1s^2 2s^2 2p^2$

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Question 2 **Essay** **0 points**
Question Give the ground-state electron configuration for fluorine (atomic number 9).
Answer $1s^2 2s^2 2p_x^2 2p_y^2 2p_z^1$ or $1s^2 2s^2 2p^5$

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Question 3 **Essay** **0 points**
Question Give the ground-state electron configuration for magnesium (atomic number 12).
Answer $1s^2 2s^2 2p^6 3s^2$

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Question 4 **Essay** **0 points**
Question How many electrons does silicon have in its valence shell?
Answer four

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Question 5 **Essay** **0 points**
Question Exhibit 1-1
 Write valid Lewis (electron-dot) structures for each formula below. Show all electrons as dots and show all non-bonding electrons.
 C_2Cl_4 tetrachloroethylene

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Question 6 **Essay** **0 points**
Question Exhibit 1-1
 Write valid Lewis (electron-dot) structures for each formula below. Show all electrons as dots and show all non-bonding electrons.
 CO_2 carbon dioxide

Answer[Add Question Here](#)[Modify](#) [Remove](#)

Question 7 **Essay** **0 points**
Question Exhibit 1-1
 Write valid Lewis (electron-dot) structures for each formula below. Show all electrons as dots and show all non-bonding electrons.
 CH_3O methanol

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Question 8 **Essay** **0 points**
Question Exhibit 1-2
 Consider the structure of urea, shown below, to answer the following question(s).



Refer to Exhibit 1-2. Fill in any non-bonding valence electrons that are missing from the line-bond structure.

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Question 9 **Essay** **0 points**

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