

Halogen Acids

Acid	pKa	Conjugate base
Inorganic Acid		Base
*HI	-10	I ⁻
*HBr	-9	Br ⁻
*HCl	-8	Cl ⁻
HF	+3.2	F ⁻

Other Strong Inorganic Acids

Acid	pKa	Conjugate base
Inorganic acid		Base
H ₂ SO ₄	-3	HSO ₄ ⁻
*H ₃ O ⁺	-2	H ₂ O
HNO ₃	+1.4	NO ₃ ⁻

Weak Inorganic Acids

Acid	pKa	Conjugate base
Inorganic acid		Base
H ₂ S	+7	HS ⁻
*NH ₄ ⁺	+9	NH ₃
*H ₂ Si ⁺	+14	SiH ₃ ⁻
*H ₂ O	+16	HO ⁻

Organic Oxyacids, Alcohols

Acid	pKa	Conjugate base
Conjugate acid		Base
*ROH ₂ ⁺	-2	ROH
*ROH	+16	RO ⁻

ac: "R" is a generic representation for various alkyl groups. The bulk of the alkyl group impact acidity but for a ballpark pKa value +16 works.

pKa Values Acids

- The smaller the value of the pKa the more acidic the acid (stronger acid)
- The more Acidic an acid is the more stable its conjugate base.
- The less Acidic an acid is the less stable its conjugate base.
- A more stable base (a weaker base) has a more acidic conjugate acid relative to a less stable base.

Acid	pKa	Conjugate Base
Carboxylic Acid		Base
acetic acid <chem>CC(=O)O</chem>	+3	<chem>CC(=O)[O-]</chem>
benzoic acid <chem>c1ccccc1C(=O)O</chem>	+4.1	<chem>c1ccccc1C(=O)[O-]</chem>
phenol <chem>c1ccccc1O</chem>	+10	<chem>c1ccccc1[O-]</chem>
4-methylphenol <chem>Cc1ccc(O)cc1</chem>	+9	<chem>Cc1ccc([O-])cc1</chem>
2-propanol <chem>CC(C)O</chem>	+16	<chem>CC(C)[O-]</chem>
1-propanol <chem>CCCO</chem>	+16	<chem>CCCO[O-]</chem>
ethanol <chem>CCO</chem>	+16	<chem>CCO[O-]</chem>
methanol <chem>CO</chem>	+16	<chem>CO[O-]</chem>
water <chem>O</chem>	+16	<chem>[O-]</chem>

Compound class	General structure	Functional group	Example
Alkanes (Chapters 3, 4)	<chem>R-H</chem>	None	<chem>CH3CH2CH2CH3</chem> Butane
Halalkanes (Chapters 6, 7)	<chem>R-X</chem> (X = F, Cl, Br, I)	<chem>-X</chem>	<chem>CH3CH2-Br</chem> Bromethane
Alcohols (Chapters 8, 9)	<chem>R-OH</chem>	<chem>-OH</chem>	<chem>CH3CH2CH2OH</chem> 1-Propanol (primary alcohol)
Ethers (Chapter 9)	<chem>R-O-R'</chem>	<chem>-O-</chem>	<chem>CH3CH2-O-CH3</chem> Methoxyethane (ethyl methyl ether)
Thiols (Chapter 9)	<chem>R-SH</chem>	<chem>-SH</chem>	<chem>CH3CH2-SH</chem> Ethaneethiol
Alkenes (Chapters 11, 12)	<chem>(R)2C=CR2</chem>	<chem>C=C</chem>	<chem>CH3CH=CH2</chem> 1-Methylpropene
Alkynes (Chapter 13)	<chem>(R)2C#C(R)2</chem>	<chem>C#C</chem>	<chem>CH3C#CCH3</chem> 2-Butyne
Aromatic compounds (Chapters 15, 16, 22)	<chem>(R)2C6H4</chem>	<chem>C6H6</chem>	<chem>c1ccccc1</chem> Methylbenzene (Toluene)
Aldehydes (Chapters 17, 18)	<chem>R-C(=O)H</chem>	<chem>-C(=O)H</chem>	<chem>CH3CHO</chem> Propanal
Ketones (Chapters 17, 18)	<chem>R-C(=O)R'</chem>	<chem>-C(=O)-</chem>	<chem>CH3COCH2CH3</chem> 3-Pentanone
Carboxylic acids (Chapters 19, 20)	<chem>R-C(=O)OH</chem>	<chem>-C(=O)OH</chem>	<chem>CH3COOH</chem> Propanoic acid
Anhydrides (Chapters 19, 20)	<chem>R-C(=O)-O-C(=O)R'</chem>	<chem>-C(=O)-O-C(=O)-</chem>	<chem>CH3COOCH2CH3</chem> Propanoic anhydride
Esters	<chem>R-C(=O)OR'</chem>	<chem>-C(=O)OR</chem>	<chem>CH3COOCH3</chem>

Amides (Chapter 19)	<chem>R-C(=O)NR2</chem>	<chem>-C(=O)NR2</chem>	<chem>CH3CH2CH2CONH2</chem> Butanamide
Nitriles (Chapter 19)	<chem>R-C#N</chem>	<chem>-C#N</chem>	<chem>CH3C#N</chem> Ethanimiditrile (Acetonitrile)
Amines	<chem>R-NR2</chem>	<chem>-NR2</chem>	<chem>(CH3)3N</chem> N,N-Dimethylamine (Trimethylamine)

n	Name	Formula
1	Methane	CH ₄
2	Ethane	CH ₃ CH ₃
3	Propane	CH ₃ CH ₂ CH ₃
4	Butane	CH ₃ CH ₂ CH ₂ CH ₃
5	Pentane	CH ₃ (CH ₂) ₃ CH ₃
6	Hexane	CH ₃ (CH ₂) ₄ CH ₃
7	Heptane	CH ₃ (CH ₂) ₅ CH ₃
8	Octane	CH ₃ (CH ₂) ₆ CH ₃
9	Nonane	CH ₃ (CH ₂) ₇ CH ₃
10	Decane	CH ₃ (CH ₂) ₈ CH ₃
11	Undecane	CH ₃ (CH ₂) ₉ CH ₃
12	Dodecane	CH ₃ (CH ₂) ₁₀ CH ₃
13	Tridecane	CH ₃ (CH ₂) ₁₁ CH ₃
14	Tetradecane	CH ₃ (CH ₂) ₁₂ CH ₃
15	Pentadecane	CH ₃ (CH ₂) ₁₃ CH ₃
16	Hexadecane	CH ₃ (CH ₂) ₁₄ CH ₃
17	Heptadecane	CH ₃ (CH ₂) ₁₅ CH ₃
18	Octadecane	CH ₃ (CH ₂) ₁₆ CH ₃
19	Nonadecane	CH ₃ (CH ₂) ₁₇ CH ₃
20	Icosane	CH ₃ (CH ₂) ₁₈ CH ₃



Organic Chemistry Exam 1 Study Guide

K. Peter C. Vollhardt, Neil E. Schore



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