

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



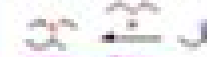
• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Electron sharing



Bonding electrons



There are two types of electronegativity

Electronegativity of atoms is



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond** is a sigma bond.

• **Two double bonds** consist of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

values, covalent bonds can be polar and nonpolar.



• If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar). For two atoms, forming a nonpolar covalent bond. Two important examples are the C-C and C-H bonds.

• If the difference in electronegativity is between 0.5 and 1.7, we have a polar covalent bond. Most common examples are atoms for carbon is bonded to N, O, F and other halogens.

• A difference of 1.7 or higher can be high that the electrons are not longer shared and a ionic bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



Electronegative atoms are color-coded (negative) and the electron density.

• The dipole moment is shown as a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• δ^- (delta minus) indicates an electron deficient atom.

• δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar (even for dipole moments are cancelled). A good example is carbon dioxide.



Cancel the vector as the other pointing point in the same way.



Two dipole moments are added (vector sum).

Bond Length and Bond Strength

The longer the bond, the weaker it is.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and a stronger bond.

Single bonds are weaker.



Double bonds are stronger.

Double bonds are stronger.



The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	472 kJ/mol	591 kJ/mol
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
C-H Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	109 pm	108 pm	106 pm
Strength	413 kJ/mol	479 kJ/mol	558 kJ/mol

Higher s character = stronger and shorter bond

Examples



sp^3 - sp^3 is a nonpolar bond of sp^3 - sp^3



sp^2 - sp^2 is a nonpolar bond of sp^2 - sp^2



sp^3 - sp^2 is a nonpolar bond of sp^3 - sp^2



sp^2 - sp is a nonpolar bond of sp^2 - sp



Double and stronger bond



sp^2 - sp is a nonpolar bond of sp^2 - sp

Organic Chemistry Study Guide

Susan McMurry



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