

Organic Chemistry Mindmap

Alkanes

- saturated C-C bond
- C_nH_{2n+2}
- UNREACTIVITY
strength and polarity of C-C bonds
- flammable
burns: $C/H_2O/CO_2 + H_2O$
soluble in organic solvents like H₂O

Hydrocarbons only contain H & C

Carboxylic Acid

- COOH functional group
- $C_nH_{2n+1}COOH$
- behaves like any other acid
reacts with metals, hydroxides, carbonates, oxides

Alcohols

- OH functional group
- $C_nH_{2n+1}OH$
- forms alcohols with reacting with alkene oxides
- test for alcohols
by oxidising them: heating, turns yellow ppt. with $K_2Cr_2O_7$

esterification
 H_2N/H_2N
catalyst

sodium carboxylate
salt drive rxn

Esters

- $C(=O)O$ functional group
- $C_nH_{2n+1}COOC_mH_{2m+1}$
- naming convention
alkyl ester from carboxylic acid
- used as flavourings

Amide

- NH_2/CH_3 functional group
- $C_nH_{2n+1}NH_2/C_nH_{2n+1}CH_3$
- fertilisers
contains N which is necessary for plant life

Alkenes

- unsaturated C=C bond
- C_nH_{2n}
- Oxidation with $KMnO_4$ or $K_2Cr_2O_7$
- addition polymerisation
only forms one polymer

electrophilic addition of halogens
 $HX + X_2$

oxidation
add reagents
 $KMnO_4$

Halogenoalkanes

- -X functional group
- $C_nH_{2n+1}X$
- S_N1 reactions
one step, forms X^- ions
- S_N2 reactions
two steps, forms tertiary halogenoalkane & tertiary carbocation
- Reaction mechanism of electrophilic addition
halogenophile breaks bond and attacks C=C bond



Aldehydes

- $C=O$ functional group
- $C_nH_{2n}CO$

oxidation
hot conc. $KMnO_4$ (primary)
 CO_2

Ketones

- $C(=O)$ functional group
- $C_nH_{2n}CO$

Markovnikov's Rule

H will bond to C atom bonded to the most alkyl groups.
Hydrogen is gay!

Carbonyls
Can be differentiated by
2,4-DNP, Fehling's solution,
Tollens' reagent



Organic Chemistry Smartarner

Bettina Heinz



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