

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equations: $K_f = 1.86^\circ\text{C}/m$, $K_b = 0.512^\circ\text{C}/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ\text{C}/m \times 5.56 \text{ mol/kg} = 10.3^\circ\text{C}$$

$\Delta T_f = 10.3^\circ\text{C}$
 $T_f = 0^\circ\text{C} - 10.3^\circ\text{C} = -10.3^\circ\text{C}$
 $\Delta T_b = 0.512^\circ\text{C}/m \times 5.56 \text{ mol/kg} = 2.85^\circ\text{C}$
 $T_b = 100^\circ\text{C} + 2.85^\circ\text{C} = 102.85^\circ\text{C}$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of urea, CH_4N_2O , and 100 g of water? Equations: $K_f = 1.86^\circ\text{C}/m$, $K_b = 0.512^\circ\text{C}/m$.
 Answer: -1.76°C

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{60 \text{ g}}}{0.100 \text{ kg}} = 16.7 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ\text{C}/m \times 16.7 \text{ mol/kg} = 31.1^\circ\text{C}$$

$\Delta T_f = 31.1^\circ\text{C}$
 $T_f = 0^\circ\text{C} - 31.1^\circ\text{C} = -31.1^\circ\text{C}$
 $\Delta T_b = 0.512^\circ\text{C}/m \times 16.7 \text{ mol/kg} = 8.55^\circ\text{C}$
 $T_b = 100^\circ\text{C} + 8.55^\circ\text{C} = 108.55^\circ\text{C}$

3. How much glucose, $C_6H_{12}O_6$, must be added to 100 g of H_2O to raise the boiling-point of the solution to 100.00°C ? Answer: 0.000 g. Assume that the solution obeys Raoult's law.

$$\Delta T_b = K_b \cdot m \quad \Delta T_b = 0^\circ\text{C} \quad m = 0$$

$$\Delta T_b = 0^\circ\text{C} = 0.512^\circ\text{C}/m \cdot m$$

$$0 = 0.512^\circ\text{C}/m \cdot m$$

$$m = 0 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} = 0$$

$$0 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg solvent}}$$

$$\text{moles solute} = 0 \text{ mol}$$

$$\text{mass solute} = \text{moles solute} \cdot \text{molar mass} = 0 \text{ mol} \cdot 180 \text{ g/mol} = 0 \text{ g}$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 1.00 g of an unknown compound has freezing-point of -1.76°C . The normal freezing-point of pure water is 0°C and the freezing-point depression constant for water is $1.86^\circ\text{C}/m$. Determine the molar mass of the compound. Answer: 40.0 g/mol.

$$\Delta T_f = K_f \cdot m \quad 1.76^\circ\text{C} = 1.86^\circ\text{C}/m \cdot m \quad m = 0.946 \text{ mol/kg}$$

$$0.946 \text{ mol/kg} = \frac{\text{mol solute}}{\text{kg solvent}}$$

$$0.946 \text{ mol/kg} = \frac{10 \text{ g}}{0.0100 \text{ kg}}$$

$$m = 0.0100 \text{ kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}}$$

$$0.946 \text{ mol/kg} = \frac{\text{moles solute}}{0.0100 \text{ kg solvent}}$$

$$\text{moles solute} = 0.00946 \text{ mol}$$

$$\text{molar mass} = \frac{\text{mass solute}}{\text{moles solute}} = \frac{10 \text{ g}}{0.00946 \text{ mol}} = 1057 \text{ g/mol}$$

$$0.00946 \text{ mol} = \frac{10 \text{ g}}{\text{molar mass}} \quad \text{molar mass} = 1057 \text{ g/mol}$$

$$\text{molar mass} = 1057 \text{ g/mol}$$

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Lingjun Ying



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