

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 10.0 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{0.177 \text{ mol}}{0.100 \text{ kg}} = 1.77 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \cdot 1.77 \text{ mol/kg} = 3.30^\circ C \text{ freezing}$$

$\Delta T_f = 1.86^\circ C/m \cdot 1.77 \text{ mol/kg}$
 $\Delta T_f = 3.30^\circ C$
 $T_f = 0^\circ C - 3.30^\circ C = -3.30^\circ C$
 $\Delta T_b = 0.512^\circ C/m \cdot 1.77 \text{ mol/kg}$
 $\Delta T_b = 0.91^\circ C$
 $T_b = 100^\circ C + 0.91^\circ C = 100.91^\circ C$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 1.00 g of urea and 100.0 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$, $M_{\text{urea}} = 60.06 \text{ g/mol}$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{0.0167 \text{ mol}}{0.100 \text{ kg}} = 0.167 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \cdot 0.167 \text{ mol/kg} = 0.311^\circ C$$

$\Delta T_f = 1.86^\circ C/m \cdot 0.167 \text{ mol/kg}$
 $\Delta T_f = 0.311^\circ C$
 $T_f = 0^\circ C - 0.311^\circ C = -0.311^\circ C$
 $\Delta T_b = 0.512^\circ C/m \cdot 0.167 \text{ mol/kg}$
 $\Delta T_b = 0.085^\circ C$
 $T_b = 100^\circ C + 0.085^\circ C = 100.085^\circ C$

3. A 100.0 mL sample of 0.100 M $NaCl(aq)$ is cooled to $-1.0^\circ C$ at which point the freezing-point of the solution is determined to be $-1.0^\circ C$. Assume that the solute does not undergo any chemical changes.

$$\Delta T_f = K_f \cdot m \quad \Delta T_f = 1.86^\circ C/m \quad m = 0.100 \text{ mol/kg}$$

$$\Delta T_f = 1.86^\circ C/m \cdot m$$

$$-1.0^\circ C = 1.86^\circ C/m \cdot m$$

$$m = 0.54 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \cdot 0.100 \text{ mol/kg}$$

$$0.54 \text{ mol/kg} = \frac{\text{moles solute}}{\text{kg solvent}} \cdot 0.100 \text{ mol/kg}$$

$$\text{moles solute} = 0.054 \text{ mol}$$

$$\text{mass solute} = \text{moles solute} \cdot M_{\text{solute}} = 0.054 \text{ mol} \cdot 58.44 \text{ g/mol} = 3.16 \text{ g}$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 1.00 g of water is cooled to $-1.0^\circ C$. The freezing-point of this solution is $-1.0^\circ C$. The normal freezing-point of pure water is $0^\circ C$. Determine the molar mass of the compound. Assume K_f is exactly $1.86^\circ C/m$.

$$\Delta T_f = K_f \cdot m \quad 1.0^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.54 \text{ mol/kg}$$

$$0.54 \text{ mol/kg} = \frac{\text{moles solute}}{\text{kg solvent}}$$

$$0.54 \text{ mol/kg} = \frac{\text{moles solute}}{0.001 \text{ kg solvent}}$$

$$m = 0.54 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \cdot 0.001 \text{ kg solvent}$$

$$0.54 \text{ mol/kg} = \frac{\text{moles solute}}{0.001 \text{ kg solvent}} \cdot 0.001 \text{ kg solvent}$$

$$\text{moles solute} = 0.00054 \text{ mol}$$

$$0.00054 \text{ mol} = \frac{10 \text{ g solute}}{M_{\text{solute}}} \quad M_{\text{solute}} = 185 \text{ g/mol}$$

$$\text{molar mass} = 185 \text{ g/mol}$$

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