

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.*

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.*

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.*

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- O C **Be** **Ne** Same Period
 - Na Rb **Fr** **H** Same Group
 - Pb** **C** Sn Si Same Group
 - Au W S **Fr** **Ne** Zn Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- O C **Be** **Ne** Same Period
- Na Rb **Fr** **H** Same Group
- Pb** **C** Sn Si Same Group
- Au W S **Fr** **Ne** Zn Challenge

Periodic Table Trends Test Answer Key

Peterson's



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Table of Contents Periodic Table Trends Test Answer Key

1. Understanding the eBook Periodic Table Trends Test Answer Key
 - The Rise of Digital Reading Periodic Table Trends Test Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Trends Test Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Trends Test Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Trends Test Answer Key
 - Personalized Recommendations
 - Periodic Table Trends Test Answer Key User Reviews and Ratings
 - Periodic Table Trends Test Answer Key and Bestseller Lists

5. Accessing Periodic Table Trends Test Answer Key Free and Paid eBooks
 - Periodic Table Trends Test Answer Key Public Domain eBooks
 - Periodic Table Trends Test Answer Key eBook Subscription Services
 - Periodic Table Trends Test Answer Key Budget-Friendly Options
6. Navigating Periodic Table Trends Test Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Trends Test Answer Key Compatibility with Devices
 - Periodic Table Trends Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Trends Test Answer Key
 - Highlighting and Note-Taking Periodic Table Trends Test Answer Key
 - Interactive Elements Periodic Table Trends Test Answer Key
8. Staying Engaged with Periodic Table Trends Test Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Trends Test Answer Key
9. Balancing eBooks and Physical Books Periodic Table Trends Test Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Trends Test Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Trends Test Answer Key
 - Setting Reading Goals Periodic Table Trends Test Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Trends Test Answer Key
 - Fact-Checking eBook Content of Periodic Table Trends Test Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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