

Unit 2: The Periodic Table Review Sheet

U2L1: PT History

- On the present day Periodic Table, elements are arranged in order of increasing atomic number
- **Groups:** vertical columns on the periodic table
 - Elements in the same group have similar chemical properties - because they have the same number of valence electrons
 - Valence electron: outermost electrons (lost # in the electron configuration)
- **Periods:** horizontal rows on the periodic table

U2L2: Parts of the Periodic Table

- Elements are classified as either metals, nonmetals, or metalloids.
 - Classification can be determined by location on the periodic table or by different properties

1

H

6

C

12.01

Atomic Mass

Symbol

Atomic Number

Electron Configuration

Selected Oxidation States

Commonly formed cations and anions

Notes: Numbers in parentheses indicate the number of isotopes or the mass of the most stable isotope.

18

He

Group

1

2

3

4

5

6

7

8

9

10

11

12

13

14

15

16

17

18

Li

Be

Na

Mg

K

Ca

Sc

Ti

V

Cr

Mn

Fe

Co

Ni

Cu

Zn

Rb

Sr

Y

Zr

Nb

Mo

Tc

Ru

Rh

Pd

Ag

Cd

Cs

Ba

La

Hf

Ta

W

Re

Os

Ir

Pt

Au

Hg

Fr

Ra

Ac

Rf

Db

Sg

Bh

Hs

Mt

Ds

Rg

Cn

Uut

Uuq

Uup

Uuh

Uus

Uuo

13

B

14

C

15

N

16

O

17

F

18

Ne

Al

Si

P

S

Cl

Ar

Ga

Ge

As

Se

Br

Kr

In

Sn

Sb

Te

I

Xe

Tl

Pb

Bi

Po

At

Rn

Periodic Table Review Pt 2

Princeton Review



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