

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave better chemical properties of elements appeared in regular intervals and whereas elements were arranged according to increasing
 - (a) density
 - (b) atomic number
 - (c) atomic mass
 - (d) electronegativity
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a column
 - (c) adding a new group
 - (d) adding a new period
5. _____ The noble gases that property of the noble gases is that they are
 - (a) unstable
 - (b) inert
 - (c) radioactive
 - (d) highly reactive
6. _____ Likewise, the first element in Group 1 has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4
 - (b) 5
 - (c) 6
 - (d) 7
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - (a) 1. How many periods are in this table?
 - (b) 2. How many elements are in this group?
 - (c) 3. What is the element symbol of this element and its atomic number?

Periodic Law Chapter 5 Test Answers

Peipei Pang



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Table of Contents Periodic Law Chapter 5 Test Answers

1. Understanding the eBook Periodic Law Chapter 5 Test Answers
 - The Rise of Digital Reading Periodic Law Chapter 5 Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Chapter 5 Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Law Chapter 5 Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Chapter 5 Test Answers
 - Personalized Recommendations
 - Periodic Law Chapter 5 Test Answers User Reviews and Ratings

- Periodic Law Chapter 5 Test Answers and Bestseller Lists
- 5. Accessing Periodic Law Chapter 5 Test Answers Free and Paid eBooks
 - Periodic Law Chapter 5 Test Answers Public Domain eBooks
 - Periodic Law Chapter 5 Test Answers eBook Subscription Services
 - Periodic Law Chapter 5 Test Answers Budget-Friendly Options
- 6. Navigating Periodic Law Chapter 5 Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
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 - Periodic Law Chapter 5 Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Law Chapter 5 Test Answers
 - Highlighting and Note-Taking Periodic Law Chapter 5 Test Answers
 - Interactive Elements Periodic Law Chapter 5 Test Answers
- 8. Staying Engaged with Periodic Law Chapter 5 Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Law Chapter 5 Test Answers
- 9. Balancing eBooks and Physical Books Periodic Law Chapter 5 Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Law Chapter 5 Test Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Law Chapter 5 Test Answers
 - Setting Reading Goals Periodic Law Chapter 5 Test Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Law Chapter 5 Test Answers
 - Fact-Checking eBook Content of Periodic Law Chapter 5 Test Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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