Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)



Atomic Rodii Size:

1. Explain atomic radius: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.

2. Using the information in figure 5-13 (pg 141) in your book, write the atomic radii size for alkali metals & period #2 on the periodic table to your right.



- 3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
- 4. Of all the elements on the periodic table, identify & circle the largest & smallest in size.
- Rank the following elements from largest to smallest in atomic radii: a) Be, Ba, Mg = Ba, Mg, Be
- 6. How can you explain the trend:

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on

in creases the attraction of the nucleus and therefore cloud close to the nucleus resulting

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Table of Contents Periodic Trends Section Review Answers

- 1. Understanding the eBook Periodic Trends Section Review Answers
 - The Rise of Digital Reading Periodic Trends Section Review Answers
 - $\circ\,$ Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Section Review Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Section Review Answers
 - Personalized Recommendations
 - Periodic Trends Section Review Answers User Reviews and Ratings
 - Periodic Trends Section Review Answers and Bestseller Lists
- 5. Accessing Periodic Trends Section Review Answers Free and Paid eBooks
 - Periodic Trends Section Review Answers Public Domain eBooks
 - Periodic Trends Section Review Answers eBook Subscription Services

- Periodic Trends Section Review Answers Budget-Friendly Options
- 6. Navigating Periodic Trends Section Review Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Periodic Trends Section Review Answers Compatibility with Devices
 - Periodic Trends Section Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Section Review Answers
 - Highlighting and Note-Taking Periodic Trends Section Review Answers
 - Interactive Elements Periodic Trends Section Review Answers
- 8. Staying Engaged with Periodic Trends Section Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Section Review Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Section Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Section Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Section Review Answers
 - Setting Reading Goals Periodic Trends Section Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Section Review Answers
 - Fact-Checking eBook Content of Periodic Trends Section Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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