

# ISOTOPES AND AVERAGE ATOMIC MASS

Name \_\_\_\_\_

Elements come in a variety of isotopes, meaning they are made up of atoms with the same atomic number but different atomic masses. These atoms differ in the number of neutrons.

The average atomic mass is the weighted average of all the isotopes of an element.

**Example:** A sample of cesium is 75%  $^{133}\text{Cs}$ , 20%  $^{132}\text{Cs}$  and 5%  $^{134}\text{Cs}$ . What is its average atomic mass?

Answer:  $.75 \times 133 = 99.75$

$.20 \times 132 = 26.4$

$.05 \times 134 = \underline{6.7}$

Total = 132.85 amu = average atomic mass

Determine the average atomic mass of the following mixtures of isotopes.

1. 80%  $^{127}\text{I}$ , 17%  $^{126}\text{I}$ , 3%  $^{128}\text{I}$

2. 50%  $^{197}\text{Au}$ , 50%  $^{198}\text{Au}$

3. 15%  $^{56}\text{Fe}$ , 85%  $^{54}\text{Fe}$

4. 99%  $^1\text{H}$ , 0.8%  $^2\text{H}$ , 0.2%  $^3\text{H}$

5. 95%  $^{14}\text{N}$ , 3%  $^{15}\text{N}$ , 2%  $^{16}\text{N}$

6. 98%  $^{12}\text{C}$ , 2%  $^{14}\text{C}$

# Periodic Table Chemistry If8766

**Michael Seilmaier**



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