



Percent Yield

45 grams of nitrogen monoxide (NO) react with an excess of sodium hydroxide (NaOH) following the balanced equation above. 6 grams of water were obtained from the reaction

What is the % yield?

Nitrogen monoxide (NO) gmm:

$$(1)(14.0) + (1)(16.0) = 30.0 \text{ g/mol}$$

$$\text{Actual Yield} = 6 \text{ g H}_2\text{O}$$

Water (H₂O) gmm:

$$(2)(1.0) + (1)(16.0) = 18.0 \text{ g/mol}$$

$$\text{Theoretical Yield} = 6.75 \text{ g H}_2\text{O}$$

$$\frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100 = 88.9\%$$

$$\frac{45 \text{ g NO} \times \frac{1 \text{ mol NO}}{30 \text{ g NO}} \times \frac{4 \text{ mol NO}}{4 \text{ mol NO}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \times 18 \text{ g H}_2\text{O}}{1} = 6.75 \text{ g H}_2\text{O}$$

Percent Yield Problems Answers

BM King



Percent Yield Problems Answers:

Unveiling the Energy of Verbal Art: An Mental Sojourn through **Percent Yield Problems Answers**

In some sort of inundated with screens and the cacophony of fast interaction, the profound power and mental resonance of verbal beauty frequently fade in to obscurity, eclipsed by the regular onslaught of sound and distractions. However, situated within the musical pages of **Percent Yield Problems Answers**, a charming work of literary beauty that impulses with fresh thoughts, lies an memorable journey waiting to be embarked upon. Written by a virtuoso wordsmith, this magical opus manuals visitors on a psychological odyssey, softly revealing the latent possible and profound influence embedded within the delicate web of language. Within the heart-wrenching expanse of the evocative evaluation, we can embark upon an introspective exploration of the book is key themes, dissect its interesting writing style, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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