

Percent Composition and Molecular Formula Worksheet Key

1. What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? **C_3H_3O mass = 55 g/mole**
2. If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula? **$C_6H_6O_2$**
3. What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? **Li_2CO_3**
4. If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula? **Li_2CO_3**

Write the molecular formulas of the following compounds:

5. A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. **$C_4O_2H_8$**
6. A compound with an empirical formula of C_4H_4O and a molar mass of 136 grams per mole. **$C_8H_8O_2$**
7. A compound with an empirical formula of $CFBrO$ and a molar mass of 254.7 grams per mole. **$C_2F_2Br_2O_2$**
8. A compound with an empirical formula of C_2H_8N and a molar mass of 46 grams per mole. **C_2H_8N**

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. **CH_2O**
10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? **$C_2H_4O_2$**
11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 9.71 g of aniline yields 6.63 g H_2O and 1.46 g N_2 , what is its empirical formula? **C_6H_7N**
12. The molar mass of aniline is 93 g/mol. What is its molecular formula? **C_6H_7N**
13. Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C_2H_5NO . **%C 40.668 %H 8.533 %N 23.713 %O 27.086**
14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? **PCl_3**
15. When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. **Hg_2O**
16. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % **C_7H_7NO**
17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?
C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %

$C_3H_7NO_3$ empirical formula

$C_3H_7NO_3$ molecular formula

Percent Composition And Molecular Formula Answers

Arun Bahl & B S Bahl



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In a digitally-driven world wherever displays reign supreme and instant conversation drowns out the subtleties of language, the profound strategies and psychological subtleties concealed within words usually get unheard. However, located within the pages of **Percent Composition And Molecular Formula Answers** a fascinating fictional value blinking with raw emotions, lies an exceptional journey waiting to be undertaken. Penned by an experienced wordsmith, this marvelous opus attracts visitors on an introspective journey, gently unraveling the veiled truths and profound affect resonating within the very material of each and every word. Within the psychological depths with this moving evaluation, we can embark upon a sincere exploration of the book is primary styles, dissect its captivating publishing design, and yield to the powerful resonance it evokes strong within the recesses of readers hearts.

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