

Organic Reactions Summary Sheet

Reactants	Products	Conditions	Classification	Example	Uses /Other
Alkanes					
+ Oxygen	Complete: CO ₂ and water Incomplete: CO and water	Little bit of heat to get it going and plenty of oxygen Little bit of heat to get it going, and not enough O ₂ , oxygen is limiting reactant	Combustion (also redox reactions)	$2 \text{C}_8\text{H}_{18} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ $2 \text{C}_8\text{H}_{18} + 9 \text{O}_2 \rightarrow 8 \text{CO} + 10 \text{H}_2\text{O}$	CO harmful to humans, unburned hydrocarbons produce smog
+ Halogen (F ₂ , Cl ₂ , Br ₂ , I ₂)	haloalkane	UV light, hν	Substitution		
Alkenes					
+ halogen (F ₂ , Cl ₂ , Br ₂ , I ₂)	Dihaloalkane		Addition		Uses: makes starting material for PVC and indication of alkenes Br ₂ dissolved in dichloromethane is used to distinguish between alkenes and alkanes. If reddish-brown color of Br ₂ disappears when added to unknown, the unknown has alkenes in it.
+ hydrogen	Alkane	Platinum or palladium catalyst	Addition (also called reduction)		Hydrogenation of oil for margarine
+ acid halide (HF, HCl, HBr, HI)	Haloalkane		Addition		Haloalkanes are useful starting materials for organic syntheses
+ water	Alcohol	Acid conditions	Addition		commercial manufacture of ethanol
Alcohols					
+ Carboxylic acid	Ester + water	Acidic solution and heat	Condensation		Esters are used in flavouring agents, plasticizers, as solvents in perfume, polyester
+ oxidizing agent (KMnO ₄ or K ₂ Cr ₂ O ₇)	Primary: Aldehyde	For incomplete oxidation: Distillation in acidic aqueous soln	Oxidation		

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DP Hallahan



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Table of Contents Organic Chemistry Reactions Summary Sheet

1. Understanding the eBook Organic Chemistry Reactions Summary Sheet
 - The Rise of Digital Reading Organic Chemistry Reactions Summary Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Reactions Summary Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry Reactions Summary Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Reactions Summary Sheet
 - Personalized Recommendations
 - Organic Chemistry Reactions Summary Sheet User Reviews and Ratings
 - Organic Chemistry Reactions Summary Sheet and Bestseller Lists
5. Accessing Organic Chemistry Reactions Summary Sheet Free and Paid eBooks
 - Organic Chemistry Reactions Summary Sheet Public Domain eBooks
 - Organic Chemistry Reactions Summary Sheet eBook Subscription Services
 - Organic Chemistry Reactions Summary Sheet Budget-Friendly Options
6. Navigating Organic Chemistry Reactions Summary Sheet eBook Formats

- ePub, PDF, MOBI, and More
- Organic Chemistry Reactions Summary Sheet Compatibility with Devices
- Organic Chemistry Reactions Summary Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry Reactions Summary Sheet
 - Highlighting and Note-Taking Organic Chemistry Reactions Summary Sheet
 - Interactive Elements Organic Chemistry Reactions Summary Sheet
- 8. Staying Engaged with Organic Chemistry Reactions Summary Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry Reactions Summary Sheet
- 9. Balancing eBooks and Physical Books Organic Chemistry Reactions Summary Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry Reactions Summary Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Organic Chemistry Reactions Summary Sheet
 - Setting Reading Goals Organic Chemistry Reactions Summary Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Organic Chemistry Reactions Summary Sheet
 - Fact-Checking eBook Content of Organic Chemistry Reactions Summary Sheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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