

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bonding and Antibonding Molecular Orbitals



Bonding molecular orbitals



There are two types of molecules

1. Polar molecules



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two multiple bonds consist of one sigma and two pi bonds.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **Polar covalent bond:** A bond between two atoms with different electronegativity values.

• **Nonpolar covalent bond:** A bond between two atoms with similar electronegativity values.

• If the difference in electronegativity is between 0.5 and 1.7, the bond is polar covalent. Examples: C-Cl, C-Br, C-I.

• If the difference in electronegativity is between 1.7 and 2.0, the bond is ionic. Examples: Na-Cl, K-Br, Li-F.

• A difference of 2.0 or higher indicates that the electrons are not shared but are lost to one atom (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The dipole moment is a vector quantity.

• The dipole moment is a vector quantity pointing from the positive end to the negative end.

• The vector dipole moment is a vector quantity pointing from the positive end to the negative end.

• **δ⁺ (delta plus)** indicates a partial positive charge.

• **δ⁻ (delta minus)** indicates a partial negative charge.

Molecular Dipole

To determine the molecular dipole, the vector sum of the bond dipoles is used.

The polarity of the molecule is a vector sum of the bond dipoles. The polarity of the molecule is a vector sum of the bond dipoles.



Draw the vector sum of the bond dipoles to find the molecular dipole.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

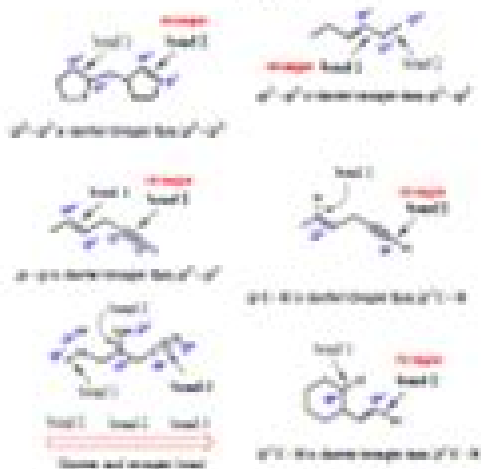


The length of the bond is related to the strength of the bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -C-C	sp ² -C-C	sp-C-C
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
C-H Bond	sp ³ -H	sp ² -H	sp-H
Length	109 pm	108 pm	106 pm
Strength	413 kJ/mol	505 kJ/mol	548 kJ/mol

Higher s character means stronger bond

Examples



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