

Halogen Acids

Acid	pKa	Conjugate base
Inorganic Acid		Base
*HI	-10	I ⁻
*HBr	-9	Br ⁻
*HCl	-8	Cl ⁻
HF	+3.2	F ⁻

Other Strong Inorganic Acids

Acid	pKa	Conjugate base
Inorganic acid		Base
H ₂ SO ₄	-3	HSO ₄ ⁻
*H ₂ O ⁺	-2	H ₂ O
HNO ₃	+1.4	NO ₃ ⁻

Weak Inorganic Acids

Acid	pKa	Conjugate base
Inorganic acid		Base
H ₂ S	+7	HS ⁻
*NH ₄ ⁺	+9	NH ₃
*H ₂ Si ⁺	+14	SiH ₃ ⁻
*H ₂ O	+16	HO ⁻

Organic Oxyacids, Alcohols

Acid	pKa	Conjugate base
Conjugate acid		Base
*ROH ₂ ⁺	-2	ROH
*ROH	+16	RO ⁻

ac: "R" is a generic representation for various alkyl groups. The bulk of the alkyl group impact acidity but for a ballpark pKa value +16 works.

pKa Values Acids

- The smaller the value of the pKa the more acidic the acid (stronger acid)
- The more Acidic an acid is the more stable its conjugate base.
- The less Acidic an acid is the less stable its conjugate base.
- A more stable base (a weaker base) has a more acidic conjugate acid relative to a less stable base.

Acid	pKa	Conjugate Base
Carboxylic Acid		Base
acetic acid	+3	acetate
benzoic acid	+4.1	benzoate
phenol	+10	phenoxide
4-methylphenol	-3	4-methylphenoxide
2-methylphenol	+5	2-methylphenoxide
3-methylphenol	+4.4	3-methylphenoxide
4-methylphenol	+2.5	4-methylphenoxide
2-methylphenol	+10.2	2-methylphenoxide
3-methylphenol	+10	3-methylphenoxide

Compound class	General structure	Functional group	Example
Alkanes (Chapters 3, 4)	R-H	None	CH ₃ CH ₂ CH ₂ CH ₃ Butane
Halalkanes (Chapters 6, 7)	R-X (X = F, Cl, Br, I)	-X	CH ₃ CH ₂ -Br Bromethane
Alcohols (Chapters 8, 9)	R-OH	-OH	CH ₃ CH ₂ -OH 1-Propanol (primary alcohol)
Ethers (Chapter 9)	R-O-R'	-O-	CH ₃ CH ₂ -O-CH ₃ Methoxyethane (ethyl methyl ether)
Thiols (Chapter 9)	R-SH	-SH	CH ₃ CH ₂ -SH Ethanesulfide
Alkenes (Chapters 11, 12)	(ROR)C=C(ROR)	C=C	CH ₃ -C=C-CH ₃ 2-Methylpropene
Alkynes (Chapter 13)	(ROR)-C≡C-(ROR)	-C≡C-	CH ₃ C≡CCH ₃ 2-Butyne
Aromatic compounds (Chapters 15, 16, 22)	(ROR) ₆ C ₆	Aromatic ring	CH ₃ -C ₆ H ₅ Methylbenzene (Toluene)
Aldehydes (Chapters 17, 18)	R-C(=O)-H	-C(=O)-H	CH ₃ CH ₂ -CHO Propanal
Ketones (Chapters 17, 18)	R-C(=O)-R'	-C(=O)-	CH ₃ CH ₂ COCH ₂ CH ₃ 3-Pentanone
Carboxylic acids (Chapters 19, 20)	R-C(=O)-OH	-C(=O)-OH	CH ₃ CH ₂ COOH Propanoic acid
Anhydrides (Chapters 19, 20)	R-C(=O)-O-C(=O)-R'	-C(=O)-O-C(=O)-	CH ₃ CH ₂ COOCH ₂ CH ₃ Propionic anhydride
Esters	R-C(=O)-OR'	-C(=O)-OR'	CH ₃ CH ₂ COOCH ₃ Methyl propanoate

Amides (Chapter 17)	R-C(=O)-NH ₂	-C(=O)-NH ₂	CH ₃ CH ₂ CONH ₂ Propanamide
Nitriles (Chapter 17)	R-C≡N	-C≡N	CH ₃ C≡N Ethanenitrile (Acetonitrile)
Amines	R-NH ₂	-NH ₂	(CH ₃) ₃ NH ⁺ N,N-Dimethylbenzylamine (Tertiary amine)

n	Name	Formula
1	Methane	CH ₄
2	Ethane	CH ₃ CH ₃
3	Propane	CH ₃ CH ₂ CH ₃
4	Butane	CH ₃ CH ₂ CH ₂ CH ₃
5	Pentane	CH ₃ (CH ₂) ₃ CH ₃
6	Hexane	CH ₃ (CH ₂) ₄ CH ₃
7	Heptane	CH ₃ (CH ₂) ₅ CH ₃
8	Octane	CH ₃ (CH ₂) ₆ CH ₃
9	Nonane	CH ₃ (CH ₂) ₇ CH ₃
10	Decane	CH ₃ (CH ₂) ₈ CH ₃
11	Undecane	CH ₃ (CH ₂) ₉ CH ₃
12	Dodecane	CH ₃ (CH ₂) ₁₀ CH ₃
13	Tridecane	CH ₃ (CH ₂) ₁₁ CH ₃
14	Tetradecane	CH ₃ (CH ₂) ₁₂ CH ₃
15	Pentadecane	CH ₃ (CH ₂) ₁₃ CH ₃
16	Hexadecane	CH ₃ (CH ₂) ₁₄ CH ₃
17	Heptadecane	CH ₃ (CH ₂) ₁₅ CH ₃
18	Octadecane	CH ₃ (CH ₂) ₁₆ CH ₃
19	Nonadecane	CH ₃ (CH ₂) ₁₇ CH ₃
20	Icosane	CH ₃ (CH ₂) ₁₈ CH ₃



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