

Name: **Answer Key**
 Date: _____
 Period: _____

P. O. Box spdf (Every electron has an address of its own)
 More Practice-Orbital Diagrams & Electron Configuration

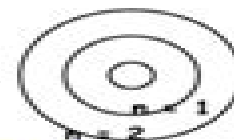


Putting It All Together...

1. The energy level $n = 1$ has only **1** sublevel, or **orbital**, called the **1s**.
 It can hold **2** electrons, just like the energy level $n = 1$ can only hold **2** electrons.



2. The energy level $n = 2$ has these sublevels, **2s, 2p_x, 2p_y, and 2p_z**. Since each orbital can only hold **2** electrons, total, all of the $n = 2$ orbitals hold **8** electrons, just like the energy level $n = 2$ can only hold **8** electrons at most.



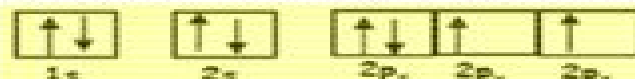
3. The energy level $n = 3$ has these sublevels, **3s, 3p_x, 3p_y, 3p_z**, as well as **five** 3d orbitals. Since each orbital can only hold 2 electrons each, total, all of the $n = 3$ orbitals hold **18** electrons, just like the energy level $n = 3$ can only hold **18** electrons at most.



More Practice with Electron Orbital Diagrams (the box thing)

Write the electron orbital diagrams (boxes with arrows) for each element below.

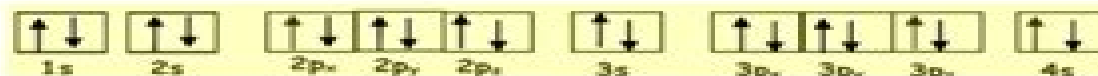
4. oxygen (O)



5. aluminum (Al)



6. calcium (Ca)



7. iron (Fe)



8. nickel (Ni)



9. arsenic (As)



10. barium (Ba)

Fill all orbitals shown for arsenic above in #9, and also include one 5s orbital filled with 2 electrons (arrows), five 4d orbitals filled with ten electrons (arrows), three 5p orbitals filled with 6 electrons (arrows), and one 6s orbital filled with 2 electrons.

Orbital Diagram Worksheet With Answers

SA Dillow



Orbital Diagram Worksheet With Answers:

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