

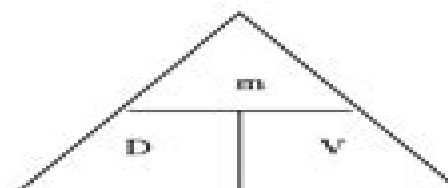
SNC 1D0
Unit: Chemistry
Density Calculations Worksheet

Name: _____

Date: _____

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

UNITS OF DENSITY
 g/cm^3 or g/mL



1. Find the unknown quantity.

a. $D = 3 \text{ g/mL}$
 $V = 100 \text{ mL}$
 $m = ?$

300 g

b. $D = ?$
 $V = 950 \text{ mL}$
 $m = 95 \text{ g}$

0.10 g/mL

c. $D = 0.5 \text{ g/cm}^3$
 $V = ?$
 $m = 20 \text{ g}$

40 cm^3

2. Find the unknown quantity (CONVERT FIRST to g or mL).

a. $D = 24 \text{ g/mL}$
 $V = 1.2 \text{ L} = \underline{\hspace{1cm}} 1200 \underline{\hspace{1cm}} \text{ mL}$
 $M = ?$

28800 g = 28.8 kg

b. $D = ?$
 $V = 100 \text{ mL}$
 $M = 1.5 \text{ kg} = \underline{\hspace{1cm}} 1500 \underline{\hspace{1cm}} \text{ g}$

15 g/mL

c. $D = ?$
 $V = 0.52 \text{ L} = \underline{\hspace{1cm}} 520 \underline{\hspace{1cm}} \text{ mL}$
 $M = 500 \text{ mg} = \underline{\hspace{1cm}} 0.5 \underline{\hspace{1cm}} \text{ g}$

0.00096 g/mL

WORD PROBLEMS

1. A block of aluminum occupies a volume of 15.0 mL and has a mass of 40.5 g. What is its density?

2.70 g/mL

2. Mercury metal is poured into a graduated cylinder that holds exactly 22.5 mL. The mercury is used to fill the cylinder has a mass of 306.0 g. From this information, calculate the density of mercury.

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Table of Contents Science Density Calculations Worksheet

1. Understanding the eBook Science Density Calculations Worksheet
 - The Rise of Digital Reading Science Density Calculations Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Science Density Calculations Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Science Density Calculations Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Science Density Calculations Worksheet
 - Personalized Recommendations
 - Science Density Calculations Worksheet User Reviews and Ratings
 - Science Density Calculations Worksheet and Bestseller Lists

5. Accessing Science Density Calculations Worksheet Free and Paid eBooks
 - Science Density Calculations Worksheet Public Domain eBooks
 - Science Density Calculations Worksheet eBook Subscription Services
 - Science Density Calculations Worksheet Budget-Friendly Options
6. Navigating Science Density Calculations Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Science Density Calculations Worksheet Compatibility with Devices
 - Science Density Calculations Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Science Density Calculations Worksheet
 - Highlighting and Note-Taking Science Density Calculations Worksheet
 - Interactive Elements Science Density Calculations Worksheet
8. Staying Engaged with Science Density Calculations Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Science Density Calculations Worksheet
9. Balancing eBooks and Physical Books Science Density Calculations Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Science Density Calculations Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Science Density Calculations Worksheet
 - Setting Reading Goals Science Density Calculations Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Science Density Calculations Worksheet
 - Fact-Checking eBook Content of Science Density Calculations Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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