

Thermochemistry Review Questions

- The process of dissolving ammonium nitrate, NH_4NO_3 , in water is an endothermic process. What is the sign of 'q'? If you were to add some ammonium nitrate to water in a flask, would you expect the flask to feel warm or cold?
- Nitric acid, a source of many nitrogen compounds, is produced from nitrogen dioxide. An old process for making nitrogen dioxide employed nitrogen and oxygen: $N_{2(g)} + 2O_{2(g)} \rightarrow 2NO_{2(g)}$. The reaction **absorbs** 56.2 kJ of heat per 2 mol of $NO_{2(g)}$ produced. Is the reaction endothermic or exothermic? What is the value of q?
- Ammonia burns in the presence of a copper catalyst to form nitrogen gas: $NH_{3(g)} + 3O_{2(g)} \rightarrow 2N_{2(g)} + 6H_2O_{(g)}$ with a $\Delta H = -1267 \text{ kJ/mol}$. What is the enthalpy change (ΔH) to burn 35.8 g of ammonia? **[Ans: -2 663.45 kJ]**
- Propane, $C_3H_{8(g)}$, is a common fuel gas. Use the following to calculate the grams of propane you would need to provide 369 kJ of heat: $C_3H_{8(g)} + 5O_{2(g)} \rightarrow 3CO_{2(g)} + 4H_2O_{(g)}$ with a $\Delta H = -2043 \text{ kJ/mol}$. **[Ans: 7.96 g]**
- Ammonia burns in the presence of a platinum catalyst to form nitric oxide: $4NH_{3(g)} + 5O_{2(g)} \rightarrow 4NO_{(g)} + 6H_2O_{(g)}$. What is the heat of reaction at constant pressure? Use the following thermochemical equations **[Ans: - 906.3 kJ]**:



- Hydrogen sulfide, H_2S , is a poisonous gas with the odour of rotten eggs. The reaction for the formation of H_2S from the element is: $H_{2(g)} + \frac{1}{8}S_{8(rhombic)} \rightarrow H_2S_{(g)}$. Use Hess' law to obtain the enthalpy change for this reaction from the following enthalpies **[Ans: - 21 kJ]**:



- The cooling effect of alcohol on the skin is due to its evaporation. Calculate the heat of vaporization (ΔH) of ethanol, $C_2H_5OH_{(l)} \rightarrow C_2H_5OH_{(g)}$. The standard enthalpy of $C_2H_5OH_{(l)}$ is - 277.7 kJ/mol and that of $C_2H_5OH_{(g)}$ is - 235.4 kJ/mol. **[Ans: 42.3 kJ]**
- Carbon disulfide is a colourless liquid. When pure, it is nearly odourless, but the commercial produce smells vile. Carbon disulfide is used in the manufacture of rayon and cellophane. The liquid burns as follows: $CS_{2(l)} + 3O_{2(g)} \rightarrow CO_{2(g)} + 2SO_{2(g)}$. Calculate the standard enthalpy change for this reaction using standard enthalpies of formation **[Ans: - 1076.6 kJ]**.
- The group IIA carbonates decompose when heated. For example: $BaCO_{3(s)} \rightarrow BaO_{(s)} + CO_{2(g)}$. Use the following enthalpies of formation and calculate the heat require to decompose 10 g of barium carbonate. The standard enthalpy of $BaCO_3$ is - 1216.3 kJ/mol. $BaO_{(s)}$ is - 942.3 kJ/mol and $CO_{2(g)}$ is - 393.5 kJ/mol. **[Ans: - 6.09 kJ]**.

Sch4u Thermochemistry Questions

**Jerome Rosenberg, Lawrence M.
Epstein, Peter Krieger**



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Commonly Asked Questions in Thermodynamics Marc J. Assael, William A. Wakeham, Anthony R. H. Goodwin, Stefan Will, Michael Stamatoudis, 2011-03-14 Have you ever had a question that keeps persisting and for which you cannot find a clear answer Is the question seemingly so simple that the problem is glossed over in most resources or skipped entirely CRC Press Taylor and Francis is pleased to introduce Commonly Asked Questions in Thermodynamics the first in a new series of books that address *THERMOCHEMISTRY* NARAYAN CHANGDER, 2024-04-08 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiz4u> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment Although the majority of students are accustomed to this MCQ format many are not well versed in it To achieve success in MCQ tests quizzes and trivia challenges one requires test taking techniques and skills in addition to subject knowledge It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations Whether you have studied the subject on your own read for pleasure or completed coursework it will assess your knowledge and prepare you for competitive exams quizzes trivia and more Experimental Thermochemistry International Union of Pure and Applied Chemistry. Commission on Thermodynamics and Thermochemistry, 1956 *Sterling Test Prep College Chemistry Practice Questions: General Chemistry Practice Questions with Detailed Explanations* Sterling Test Prep, 2018-01-05 College Chemistry bestseller General chemistry practice questions with detailed explanations that cover all inorganic chemistry topics Electronic and atomic structure of matter Periodic table Chemical bonding States of matter gases liquids solids Solution chemistry Acids and bases Stoichiometry Equilibrium and reaction rates Thermochemistry Electrochemistry This book provides 1 200 practice questions that test your knowledge of inorganic chemistry topics The detailed explanations discuss why the answer is correct and more importantly why another answer that may have seemed correct is the wrong choice The explanations include the foundations and details of important science topics needed to answer related questions on your tests By reading these explanations carefully and understanding how they apply to solving the question you will learn important chemistry concepts and the relationships between them This will prepare you for your college chemistry exams and will significantly improve your scores All the questions in this book are prepared by chemistry instructors with years of

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Experimental Thermochemistry Frederick D. Rossini,H. A. Skinner,1962

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Questions in Chemistry A. L. H. SMITH, 1972

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