

CHEMISTRY FORM 4 SCHEMES OF WORK TERM 1,2,3

| FORM FOUR CHEMISTRY | | | | | | | |
|---------------------|-------|--|---|--|--|-------------------------------|---------|
| SCHEME OF WORK | | | | TERM ONE | | | |
| WK NO | LJ NO | TOPIC / SUBTOPIC | LESSON / SPECIFIC OBJECTIVES | TEACHING / LEARNING ACTIVITIES | MATERIALS / RESOURCES | REFERENCES | REMARKS |
| 1 | | ACIDS, BASES AND SALTS. | <u><i>By the end of the lesson, the learner should be able to</i></u> | | | | |
| | 1. | Strength of acids. | Define an acid in terms of hydrogen ions. | Class experiments: investigate reactions of magnesium and zinc carbonate with different acids. Make and record observations in tabular form. | Magnesium strip, zinc carbonate, 2M HCl, 2M H ₂ SO ₄ , 2M ethanoic acid. | <i>K.L.B. BK IV Pages 1-4</i> | |
| | 2 | Acids in aqueous form. | Explain strength of acids in aqueous form in terms of number of hydrogen ions present. | Make deductions from the observations. Write relevant chemical equations and ionic equations. Detailed discussion leading to the definition of an acid and explanation of strength of an acid. | | | |
| | 3. | pH values of acids. | Determine strength of acids using pH values. | Q/A: review determination of strength of acids using a litmus paper and pH scale. Class / group experiments: record colour of universal indicator in 2M HCl and 2M ethanoic acid. | Universal indicator, 2M HCl, 2M ethanoic acid, dry cells, carbon electrodes, milli-ammmeters, wires, switches etc. | <i>K.L.B. BK IV Pages 4-6</i> | |
| | 4 | Electrical conductivities of aqueous acids. | Determine strengths of acids by comparing their electrical conductivities. Classify acids as either strong or weak in terms of partial dissociations in aqueous solutions. | Set up voltmeters of 2M HCl and 2M ethanoic acid in turn. Record ammeter of current . Discuss the observations. Write corresponding ionic equations. | | | |
| | 5 | Definition of a base in terms of hydroxide ions. | Define a base in terms of hydroxide ions. | Teacher demonstration: Dissolve calcium hydroxide in water. Carry out litmus test on the resulting solution. Discuss the results; hence define a base in terms of hydroxide ions. | Red litmus paper, calcium hydroxide solid. | <i>K.L.B. BK IV Pages 6-7</i> | |

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