



# Schematic Diagram Of Automatic Lead Battery Charger

**AW Chickering**



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 $pK_b$  of the weak acid or base being titrated a solution ph as a function of the volume of 1.00 M NaOH added to 10.00 ml of  
1.00 M solutions of weak acids with the indicated  $pK_a$  values

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the word titration comes from the latin word titulus which means inscription or title

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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should  
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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the  
ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $ph = -\log 0.100 = 1.000$  ph  $-\log 0.100 = 1.000$  b titrant volume 12.50 ml

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reacting it with a standard solution one type of titration uses a neutralization reaction in which an acid and a base react to  
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web at the equivalence point in an acid base titration moles of base = moles of acid and the solution only contains salt and  
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web expert answer 100% 2 ratings the example is based upon Henderson-Hasselbalch equation the equation gives relation

between  $pH$  and  $pK_a$  of weak acid the equation is as  $pH = pK_a + \log \frac{[salt]}{[acid]}$  at equivalence point all a view the full answer transcribed image text

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web titration curves source openstax chemistry atoms first 2e strong acid strong base titrations when titrating a strong acid such as hydrochloric acid with sodium hydroxide you are reacting the HCl directly with NaOH  $HCl(aq) + NaOH(aq) \rightarrow H_2O(l) + NaCl(aq)$

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Volume of NaOH added (mL)	pH
4.6	4.6
4.6	5.9
5.9	5.9

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web background plot of the ph of a solution against the volume of titrant added is called a titration curve the ph can be measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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